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Determination of the Reactivity Ratios of Methyl Acrylate with the Vinyl Esters Vinyl Acetate, Vinyl 2,2-Dimethyl-Propanoate, and Vinyl 2-Ethylhexanoate

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SYNOPSIS

The course of composition drift in copolymerization reactions is determined by reactivity ratios of the contributing monomers. Since polymer properties are directly correlated with the resulting chemical composition distribution, reactivity ratios are of paramount importance. Furthermore, obtaining correct reactivity ratios is a prerequisite for good model predictions. For vinyl acetate (VAc), vinyl 2,2-dimethyl-propanoate also known as vinyl pivalate (VPV), and vinyl 2-ethylhexanoate (V2EH), the reactivity ratios with methyl acrylate (MA) have been determined by means of low conversion bulk polymerization. The mol fraction of MA in the resulting copolymer was determined by 'H-NMR. Nonlinear optimization on the thus-obtained monomer feed–copolymer composition data resulted in the following sets of reactivity ratios: $r_{MA} = 6.9 \pm 1.4$ and $r_{VAc} = 0.013 \pm 0.02$; $r_{MA} = 5.5 \pm 1.2$ and $r_{VPV} = 0.017 \pm 0.035$; $r_{MA} = 6.9 \pm 2.7$ and $r_{V2EH} = 0.093 \pm 0.23$. As a result of the similar and overlapping reactivity data of the three methyl acrylate–vinyl ester monomer systems, for practical purposes these data can be described with one set of reactivity data. Nonlinear optimization of all monomer feed–copolymer composition data together resulted in $r_{MA} = 6.1 \pm 0.6$ and $r_{VEst} = 0.0087 \pm 0.023$. © 1994 John Wiley & Sons, Inc.

Keywords: bulk • copolymerization • reactivity ratios • methyl acrylate • vinyl esters

INTRODUCTION

Vinyl acetate copolymers are widely used in interior architectural coatings. Due to the poor hydrolytic stability of these copolymers, their use as exterior coatings is limited. Vinyl ester monomers of the C9–C11 versatic acids, as produced by Shell, have been available in Europe for about 25 years. As a result of recently developed large scale transvinylation methods, new vinyl ester monomers have been produced by the Union Carbide Corporation. Due to neighboring group steric effects, the use of these monomers in emulsion copolymerization results in improved hydrolytic stability and water resistance when compared with vinyl acetate. Furthermore, glass transition temperatures of (co)polymers will strongly depend on the vinyl ester used. In this way copolymers can be designed over a wide range of glass transition temperatures.

Next to resistance against hydrolysis and glass transition temperatures, the product properties are also determined by the heterogeneity of the copolymer. Predictions of composition drift are, therefore, very important. With the reactivity ratios at hand, predictions of the chemical composition distribution can be performed for copolymerization reactions in both homogeneous (bulk and solution) and heterogeneous (emulsion) systems. Although only small differences in reactivity ratios have been observed in vinyl acetate–vinyl esters copolymerization reactions ($r_{VAc} \approx r_{VEst}$), composition drift occurring in copolymerization reactions of, for instance, an acrylic monomer with vinyl acetate may be affected by replacing vinyl acetate by another vinyl ester. Since acrylic polymers are often used in various applications, it is important to know whether
or not the reactivity ratio of MA with vinyl esters can be approximated by the reactivity ratio of vinyl acetate, often already known. For this reason the reactivity ratios of methyl acrylate–vinyl esters have been determined for vinyl acetate (VAc), vinyl 2,2-dimethyl-propanoate (VPV), and vinyl 2-ethylhexanoate (V2EH). These monomers have been selected since they do not consist of isomeric mixtures.

The general structure of the MA-VEst copolymers is given by:

In this structural formula the left hand group \( x \) represents the MA units and the right hand group \( y \) represents the vinyl ester units in the copolymer. The R-group in the structural formula stands for a CH\(_3\), C\(_4\)H\(_9\), and C\(_7\)H\(_{15}\) group representing the monomers VAc, VPV, and V2EH, respectively.

For MA-VAc copolymerizations the reactivity ratios have been determined before by several investigators using different copolymer analysis methods, as polymer hydrolysis followed by acetic acid determinations,\(^5\) infrared spectroscopy,\(^6\) interferometry,\(^7\) and \(^1\)H-NMR.\(^8\) However, the monomer feed–copolymer composition data thus obtained sometimes lack accuracy.\(^5\) Furthermore, the reactivity ratios have been calculated by traditional linearization techniques\(^9\),\(^10\),\(^11\) of the instantaneous copolymer equation\(^12\),\(^13\) given by:

\[
F_i = \frac{r_i f_i^2 + f_i f_j}{r_i f_i^2 + 2 f_i f_j + r_j f_j^2}
\]

(1)

where \( F_i \) stands for the copolymer composition in mol fraction of monomer, \( f_i \), and \( f_j \) are the molar fractions of monomers \( i \) and \( j \) in the feed, and \( r_i \) and \( r_j \) are the reactivity ratios of monomers \( i \) and \( j \) in the selected monomer combination \( i-j \). Recently, more accurate nonlinear optimization techniques have been developed to determine reactivity ratios.\(^14\),\(^15\) In this article the reactivity ratios are determined by the nonlinear optimization technique described by Dubé et al.\(^15\) taking into account the estimated experimental error in both monomer feed and copolymer composition.

A typical 400 MHz \(^1\)H-NMR spectrum of a low conversion bulk MA-V2EH copolymer is given in Figure 1. The total peak area represented by \( A \) in Figure 1 is generated as a result of the resonance of the V2EH proton marked by \('*'\) in the above given structural formula of the copolymer. The peak area \( B \) is generated by the three methyl group protons of MA in the copolymer. This methyl group is indicated by \('**'\) in the structural formula. The peak area generated by all other protons is represented by \( C \) in Figure 1. Note that the number of protons in the side group (and, therefore, the total peak area of \( C \) as compared with \( A \) and \( B \)) will change with each selected vinyl ester. Calculation of the average copolymer composition of all MA-VEst copolymers (mol fraction MA: \( F_{MA} \)) can be determined from the peak areas \( A \) and \( B \) in the \(^1\)H-NMR spectra by using the following relationship:

\[
F_{MA} = \frac{B}{3A + B}
\]

(2)

**EXPERIMENTAL**

**Materials**

The following materials were used for the low conversion bulk copolymerization reactions: reagent grade methyl acrylate (MA, Janssen Chimica), vinyl acetate (VAc, Janssen Chemica), vinyl 2,2-dimethyl-propanoate (VPV, product names VEOVA-5 from Shell and Vynate NEO-5 from Union Carbide Corporation), and vinyl 2-ethylhexanoate (V2EH,
product name vynate 2EH, Union Carbide Corporation) as monomers, and 2,2'-azobis(2-methylpro- 
pionitrile) (AIBN, Janssen Chimica, Tilburg, The Netherlands) as initiator. Before use, the monomers 
were distilled under reduced pressure in order to remove inhibitor. The middle fraction was cut and 
stored at 4°C.

Determination of the Reactivity Ratios from Bulk Polymerizations

Bulk copolymerizations were carried out in 20-mL bottles thermostated at 50°C and mixed with a 
magnetic stirrer. The reaction mixture consisted of 20 g of monomer with different monomer mol frac- 
tions ranging from 10-90 mol % MA and 0.1 g AIBN as initiator. The reactions were stopped at low con- 
version (<3%) in order to prevent composition drift, by pouring the total reaction mixture in heptane. 
To determine the copolymer composition of the re- 
sulting copolymer, a 400 MHz $^1$H-NMR spectrum 
(Bruker AM 400) was taken at 298 K using CDCls 
as solvent and locking agent. From the relative peak 
areas in this spectrum the copolymer composition 
could be determined. The resulting monomer feed 
fraction–copolymer composition data are shown in 
Table I.

RESULTS AND DISCUSSION

Reactivity Ratios

Predictions of the course of copolymerization re- 
actions as a function of conversion can be performed 
only in those cases where the reactivity ratios are 
known accurately. For the monomer combination 
MA-VAc the reactivity ratios have been determined 
before by Kulkarni et al.$^5$ resulting in the following 
reactivity ratios: $r_{MA} = 6.3 \pm 0.4$ and $r_{VAc} = 0.031 
\pm 0.006$. However, these reactivity ratios have been 
determined by linearization of the instantaneous 
copolymer equation. Recently, better nonlinear op- 
timization techniques have been developed,$^{14,15}$ 
making more realistic determinations of the react- 
ivity ratios possible by taking into account exper- 
imental errors in the monomer feed and copolymer 
composition. The results of newly obtained mono- 
mer feed–copolymer composition data for MA-VAc 
are given in Table I.

For the monomer combinations MA-VPV and 
MA-V2EH no reactivity ratios were found in the 
literature. In order to see whether or not the reac- 
tivity ratios for the investigated MA-VEst systems 
are similar, low conversion bulk polymerizations 
have to be performed in order to actually determine 
the reactivity ratios. These reactivity values were 
determined by nonlinear optimization$^{15}$ of the 
monomer feed–copolymer composition data sum- 
marized in Table I. When calculating reactivity ra- 
tios with this nonlinear optimization method, errors 
in both the monomer feed composition (estimated 
to be 0.1%) and the copolymer composition (esti- 
"mate to be 3%) are taken into account.

The reactivity ratios resulting from nonlinear 
optimization of the monomer feed–copolymer com- 
position data (Table I) are given in Figure 2, to- 
gether with the 95% reliability intervals and in Table 
II. It should be mentioned that the large reliability 
interval on the reactivity ratio of V2EH in the MA- 
V2EH monomer combination is a result of the lack 
of monomer feed–copolymer composition data at low 
MA feed ratios. Bulk polymerizations of MA-V2EH, 
performed at low MA feed ratios, were extremely

?Copolymer composition ($F_{\text{MA}}$) data (both in monomer mol fractions) of the monomer system MA-VAc, MA-VPV, and MA-V2EH obtained by $^1$H-NMR of low conversion bulk polymerizations products.
sensitive to inhibition. For this reason no reactions at low MA feed ratios could be performed. From the results in Table II and Figure 2 it can be concluded that no significant differences are found between the reactivity ratios of the three monomer systems. Therefore, for practical purposes they can be estimated with one set of reactivity ratios. This pair of overall reactivity ratios was obtained by nonlinear optimization of all monomer feed–copolymer composition data of the three MA–VEst systems together. Due to the increased number of data points, the 95% reliability interval and the error bars on the reactivity ratios have decreased resulting in the following reactivity ratios:

\[ r_{MA} = 6.06 \pm 0.6 \text{ and } r_{VEst} = 0.0087 \pm 0.023. \]

This result is in acceptable agreement with the reactivity ratios determined by Kulkarni et al.\(^8\) for MA–VAc \((r_{MA} = 6.3 \pm 0.4 \text{ and } r_{VAc} = 0.031 \pm 0.006)\).

When using the overall set of reactivity ratios, the monomer feed–copolymer composition relationship can be described theoretically using the instantaneous copolymer equation \([\text{Eq. (1)}]\). Comparing experimental results with the theoretical prediction of the instantaneous copolymer equation using the reactivity ratios \(r_{MA} = 6.06 \text{ and } r_{VEst} = 0.0087\) gives good agreement, as can be seen in Figure 3. From these results it can be concluded that all three monomer combinations indeed can be described with one set of reactivity ratios.

An advantage of the approximately equal reactivity ratios for VAc, VPV, and V2EH in MA–VEst monomer systems is that these systems can be used as a tool in studying the important effect of the monomer solubility in water on the course of emulsion copolymerization reactions of MA–VEst, as a function of the monomer to water ratio. We will report on these aspects in a separate article.

### Table II. Reactivity Ratios of MA–VEst

Monomer Combinations as Determined from Low Conversion Bulk Polymerizations at 50°C

<table>
<thead>
<tr>
<th>Vinyl Ester</th>
<th>MA</th>
<th>VEst</th>
</tr>
</thead>
<tbody>
<tr>
<td>Vinyl acetate</td>
<td>6.90 ± 1.4</td>
<td>0.013 ± 0.02</td>
</tr>
<tr>
<td>Vinyl 2,2-dimethyl-propanoate</td>
<td>5.5 ± 1.2</td>
<td>0.017 ± 0.035</td>
</tr>
<tr>
<td>Vinyl 2-ethylhexanoate</td>
<td>6.9 ± 2.7</td>
<td>0.093 ± 0.23</td>
</tr>
<tr>
<td>Overall vinyl esters</td>
<td>6.1 ± 0.6</td>
<td>0.0087 ± 0.023</td>
</tr>
</tbody>
</table>

### CONCLUSIONS

Reactivity ratios of methyl acrylate with the vinyl esters vinyl acetate, vinyl 2,2-dimethyl-propanoate, and vinyl 2-ethylhexanoate have been determined by a nonlinear optimization technique applied to the monomer feed–copolymer composition data obtained from low conversion bulk polymerization reactions. This resulted in the following reactivity ratios:
MA-VAc: $r_{\text{MA}} = 6.9 \pm 1.4$ and $r_{\text{VAc}} = 0.013 \pm 0.03$;

MA-VPV: $r_{\text{MA}} = 5.5 \pm 1.2$ and $r_{\text{VAc}} = 0.017 \pm 0.035$;

MA-V2EH: $r_{\text{MA}} = 6.9 \pm 2.7$ and $r_{\text{VAc}} = 0.093 \pm 0.23$.

On the basis of the above shown overlapping reactivity ratios it was concluded that the reactivity of the three MA-VEst monomer systems can be described with one set of reactivity ratios. Nonlinear optimization of all monomer feed–copolymer composition data of the MA-VEst systems resulted in the following overall reactivity data: $r_{\text{MA}} = 6.1 \pm 0.6$ and $r_{\text{VEst}} = 0.0087 \pm 0.023$ (50°C).

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