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A NOVEL SYNTHESIS OF VINYL ESTERS FROM VINYLVERSATE-10

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Abstract—Vinylversate-10 (VV10) has successfully been used to synthesise a large number of lower vinyl esters by transvinylaion in presence of mercuric acetate and sulfuric acid. The synthesis of vinylhalo esters proceeds with more difficulty. It has been observed that neither Hg(OAc)₂ nor H₂SO₄ alone is capable of initiating the transvinylaion. Furthermore, it has been found that a molar ratio 2:1 of VV10 to carboxylic acid is sufficient to drive the reaction to the right by continuous distillation of the vinyl ester formed, and as a result a high yield of vinyl ester is obtained. A mechanism for this reaction and for the formation of side products has been proposed.

INTRODUCTION

Although the synthesis of vinyl esters has been known since 1912, most of the methods are useful in specific cases only, and as laboratory methods the yields are poor.

The classical method is the direct vinylation using acetylene and carboxylic acid. This reaction has only been used to produce the esters on an industrial scale. In spite of engineering problems (e.g. corrosion) it appears that for vinyl acetate the ethylene based process may replace the acetylene process. Most of the vinyl esters other than vinyl acetate, vinyl propionate and vinyl versatates, used today, are not produced on a large scale. The transvinylaion or vinyl interchange method is commonly and conveniently used in the laboratory, but also in industry, especially when purification by distillation is difficult.

The transvinylaion reaction using vinyl acetate in presence of catalytic amounts of mercuric salts was first described by Herrmann and Haehnel in 1941 and by Toussaint and Mc-Dowell in 1943. The main side reaction, the formation of ethylidene diacetate, was controlled by running the reaction at a low temperature for a long time. Chatterjee et al. used HF in combination with mercuric acetate to shorten the reaction time and improved the overall yield. Use of PdCl₂ instead of mercuric salts also gave higher yields. Warson et al. described the preparation of vinyl esters by vinyl interchange of C₇-C₂₆ acids with vinyl acetate using mercuric salts of aromatic sulfonic acids as catalyst.

In many cases, the synthesis of vinyl esters from vinyl acetate by the transvinylaion method requires refluxing for a long time and the formation of ethylidene diacetate is very high. Consequently the yield of vinyl ester is very low.

Shell Chemical Company produces a number of vinyl esters of branched carboxylic acids. These vinyl esters (called vinyl versatates) are easily available, cheap and have relatively high boiling points. A new procedure has, therefore, been obtained, which is applicable to the synthesis of lower vinyl esters starting from vinyl versatate-10 (VV10). Advantage was taken of the fact that the desired vinyl ester is directly distilled from a solution containing catalyst, vinylating agent and carboxylic acid. This is feasible since vinyl esters generally have b.p.s about 40-50⁰ below those of the acids from which they are derived. Thus changing the vinylating agent from vinyl acetate to VV10 improves the overall yield and produces vinyl esters easily.

EXPERIMENTAL

General experimental procedure. The transvinylaion reaction was carried out in a 3-necked one litre flask fitted with a glass stirrer, N₂ inlet tube and a distillation column. A mixture of two moles of VV10 and one mole of carboxylic acid together with 2% mercuric acetate by weight of acid, 0.005 mole of H₂SO₄ and 2-4 g of p-t-butylcatechol (inhibitor) was introduced into the flask. Before heating and during the reaction N₂ gas was passed through the mixture and the contents was stirred. The temp was controlled by using a constant temp silicone oil bath. The vinyl ester formed was collected directly from the flask during the reaction by continuous distillation.

It has been observed by GLC in every synthesis that the product obtained by continuous distillation contained some acetaldehyde and vinyl acetate. Traces of metallic mercury were also found inside the cooler. Finally, the product obtained was distilled once more using a fractionating column. The purity of the synthesised vinyl esters was checked by b.p, refractive index, IR spectrum and GLC. The reaction conditions, percentage conversion and physical constants of the vinyl esters are shown in Table I.

DISCUSSION

It will be evident from the results shown in Table 1 that the yields of vinyl esters are very high as compared with the literature values, where vinyl acetate was used. Heating for 3 hr was also found to be sufficient to obtain vinyl esters in good yields from VV10 instead of 8-12 or more hr when starting from vinyl acetate. This is possible because of the fact that the desired vinyl ester was continuously distilled from a solution of catalyst, a high boiling vinylating agent (VV10) and carboxylic acid, thus shifting the equilibrium to the right.

$$VV10 + RCOOH \rightarrow RCOOH = CH_2 + V10\text{acid}$$
Table 1. Synthesis of vinylesters from vinylversatate (VV10)

<table>
<thead>
<tr>
<th>Name of acid</th>
<th>Acid mole</th>
<th>VV10 mole</th>
<th>Mercuric acid Acetate g</th>
<th>Sulfuric acid mole</th>
<th>Temp °C</th>
<th>Time hour</th>
<th>% Conversion based on acid</th>
<th>% Yield*</th>
<th>B.p of ester °C</th>
<th>nD</th>
</tr>
</thead>
<tbody>
<tr>
<td>Formic a</td>
<td>1</td>
<td>2</td>
<td>1.4</td>
<td>0.005</td>
<td>95</td>
<td>4</td>
<td>35</td>
<td>60</td>
<td>44-46</td>
<td>1.3420</td>
</tr>
<tr>
<td>Acetic a</td>
<td>1</td>
<td>2</td>
<td>1.5</td>
<td>0.005</td>
<td>90</td>
<td>3</td>
<td>70</td>
<td>90</td>
<td>72-73</td>
<td>1.39420</td>
</tr>
<tr>
<td>Propionic a</td>
<td>1</td>
<td>2</td>
<td>1.5</td>
<td>0.005</td>
<td>120</td>
<td>3</td>
<td>80</td>
<td>90</td>
<td>93-5-95</td>
<td>1.40050</td>
</tr>
<tr>
<td>Butyric a</td>
<td>1</td>
<td>2</td>
<td>1.6</td>
<td>0.005</td>
<td>140</td>
<td>3</td>
<td>75</td>
<td>80</td>
<td>116-118</td>
<td>1.41130</td>
</tr>
<tr>
<td>Isobutyric b</td>
<td>1</td>
<td>2</td>
<td>1.8</td>
<td>0.005</td>
<td>120</td>
<td>3</td>
<td>70</td>
<td>75</td>
<td>103-5-105</td>
<td>1.40560</td>
</tr>
<tr>
<td>Chloroacetic a</td>
<td>1</td>
<td>2</td>
<td>2.0</td>
<td>0.006</td>
<td>150</td>
<td>6</td>
<td>60</td>
<td>70</td>
<td>135-136</td>
<td>1.44440</td>
</tr>
<tr>
<td>Bromoacetic a</td>
<td>1</td>
<td>2</td>
<td>2.5</td>
<td>0.006</td>
<td>160</td>
<td>3</td>
<td>25</td>
<td>75</td>
<td>145</td>
<td>1.47500</td>
</tr>
<tr>
<td>Dichloroacetic b</td>
<td>2</td>
<td>4</td>
<td>6.0</td>
<td>0.012</td>
<td>125</td>
<td>7</td>
<td>35</td>
<td>45</td>
<td>48-52</td>
<td>1.45395</td>
</tr>
<tr>
<td>Trichloroacetic b</td>
<td>1</td>
<td>2</td>
<td>3.0</td>
<td>0.006</td>
<td>150</td>
<td>5</td>
<td>No vinyl interchange reaction</td>
<td>n/a</td>
<td>n/a</td>
<td>n/a</td>
</tr>
<tr>
<td>Trifluoroacetic b</td>
<td>1</td>
<td>2</td>
<td>3.0</td>
<td>0.006</td>
<td>70</td>
<td>5</td>
<td>No vinyl interchange reaction</td>
<td>n/a</td>
<td>n/a</td>
<td>n/a</td>
</tr>
</tbody>
</table>

*% Yield calculated from % conversion on mole basis after purification; aMerck; bFluka; cRiedel-de Haén.

Moreover, further reaction of the newly formed vinyl ester with the added carboxylic acid to form acylyl is also minimised.

Considering the mechanism reported by Slinckx and Smets 11 and Hopff and Osman 12,13 for the transvinylisation using vinyl acetate, we propose the following scheme for the vinyl interchange with VV10.

This mechanism is in agreement with our experimental results. The catalyst acts as follows:

\[ H(OAc)_2 + H^+ \rightarrow HgOAc + HOAc \]

The cation formed reacts with VV10 giving a complex,

\[ C_6H_{10}COOCH\equiv CH_2 + HgOAc \]

This complex ion (1) then reacts with an anion derived from the added acid giving an intermediate compound (2).

\[ [C_6H_{10}COOCH\equiv CH_2]^- + RCOO \]

\[ C_6H_{10}COOCHCH_2HgOAc \]

\[ RCOO \]

\[ C_6H_{10}COOCHCH_2HgOAc \]

\[ RCOO \]

\[ RCOOCH\equiv CH_2 + HgOAc + C_6H_{10}COOH \]
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The equilibrium (a) depends on the nature of R in such a way that a more electron donating group R will shift the equilibrium backward.

Though R has influence on the equilibrium (a), it also determines the elimination equilibrium (b) which is the rate determining step. This elimination step will be faster when R is an electron donating group, i.e. the vinyl interchange will be favoured with weaker acids (RCOOH). Thus going from \(-\text{CX}\) via \(-\text{CX}_2\) to \(-\text{CX}_3\) (X = halogen), the group R becomes more electron withdrawing, so the chance of occurrence of the elimination reaction (b) becomes increasingly smaller. This scheme is consistent with our experimental results, monochloro- and monobromoacetic acid react rather slowly, dichloroacetic acid reacts still slower, trichloro- and trifluoroacetic acid do not give vinyl esters at all.

Moreover, as reported by Sandler\(^1\), the failure of the synthesis of vinyltrichloro- and trifluoro-esters from vinylacetate by the transvinylolation method may be due to the formation of acylal directly:

\[
\begin{align*}
\text{CH}_2\text{C}=\text{O} & \quad \text{H}_2\text{O} \\
\text{CH}_2\text{CHO} + \text{CH}_2\text{COOH} \rightarrow \text{CH}_2=\text{C} & \quad \text{CH}_2=\text{CHCOOH} \\
\text{CH}_2=\text{C} & \quad \text{CH}_2=\text{O}
\end{align*}
\]

Any acyl formed during the reaction may also decompose to give anhydride and acetaldehyde.

\[
\begin{align*}
\text{CH}_2=\text{C} & \quad \text{O} \\
\text{CH}_2 & \quad \text{CH}_2=\text{O}
\end{align*}
\]

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