Cation distribution in calcium-strontium-hydroxyapatites

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Table 1. Eigenvalues and eigenvectors of $^3H_4$ term in the crystal field of $C_2\bar{v}$ symmetry calculated with values listed in Fig. 2. $(J_r)$ represents $-J_2$.

<table>
<thead>
<tr>
<th>Eigenvalue ($cm^{-1}$)</th>
<th>Eigen vector</th>
</tr>
</thead>
<tbody>
<tr>
<td>4784</td>
<td>$0.0171(4\rangle - 0.4092(2\rangle + 0.8153(0\rangle$</td>
</tr>
<tr>
<td>4780</td>
<td>$-0.1454(3\rangle - 0.6920(1\rangle + 0.1267(1\rangle$</td>
</tr>
<tr>
<td>1578</td>
<td>$-0.1267(4\rangle - 0.6957(2\rangle$</td>
</tr>
<tr>
<td>1560</td>
<td>$-0.4466(3\rangle - 0.5482(1\rangle$</td>
</tr>
<tr>
<td>-1135</td>
<td>$-0.1804(4\rangle + 0.5551(2\rangle + 0.5646(0\rangle$</td>
</tr>
<tr>
<td>-1107</td>
<td>$0.6920(3\rangle + 0.1454(1\rangle + 0.6957(1\rangle$</td>
</tr>
<tr>
<td>-2162</td>
<td>$0.5482(4\rangle + 0.1267(2\rangle$</td>
</tr>
<tr>
<td>-4040</td>
<td>$0.6957(4\rangle + 0.1567(2\rangle + 0.1287(0\rangle$</td>
</tr>
<tr>
<td>-4056</td>
<td>$0.6835(4\rangle + 0.1567(2\rangle + 0.1287(0\rangle$</td>
</tr>
</tbody>
</table>

The calculation of the magnetic susceptibility from the energy splitting behavior thus obtained has been done with use of the Van Vleck formula to the first excited state, including an additional open variable $Q$ introduced by Amberger et al.[10]. As shown in Fig. 1, a good agreement with the experimental result up to 170°K is obtained, when $R_4$ and $Q$ are taken to be 2.35 Å and 0.829, respectively, (2θ = 67.2°). The eigenvalues and the eigenvectors are listed in Table 1. The value of $Q$ is intermediate between that of uranocene ($Q = 0.94$)[11] and that of ($\text{CsH}_0\text{U(IV)}$ ($Q = 0.707$)[10]. The distance $R_4$ and the angle $\theta_4$ seems to be reasonable values considering that the U-N distance of U(NCS)$_5\text{CsH}_8\text{N}$ is 2.38 Å[12] and the angle of $\text{O}^{\text{HF}}$-$\text{U}$-$\text{O}^{\text{HF}}$ is 67.2°[5]. The small deviation from the experimental value above about 170°K may be explained by taking into account higher excited states.

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References

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Contradictory data about the effect of Sr$^{2+}$ ions in drinking water on dental caries[1,2] necessitate a reinvestigation of Ca$^{2+}$ and Sr$^{2+}$ containing phosphates. Especially those with the apatite structure might be of interest as calcium hydroxyapatite (CaOHA) has been reported to form a continuous series of solid solutions with strontium hydroxyapatites (SrOHA)[3,4], whereas strontium analogues of octocalcium phosphate Ca$\text{H}$PO$_4$.2H$_2$O do not exist[5,6]. Moreover, SrOHA can incorporate large amounts of carbonate like the biological apatites[6] which are derived from CaOHA as the prototype[7].

In this study pure CaOHA and SrOHA were prepared by titrating a boiling slurry of either calcium- or strontiumhydroxyde with phosphoric acid[8]. Solid solutions of the formula

$$\text{Ca}_{10-x}\text{Sr}_x\text{PO}_{4x}(x\text{OH})_2 \quad (1)$$

at $x = 1, 2, 4, 5, 6, 8$ and 9 were prepared by solid state reaction of the respective ternary apatites at 1200°C in a stream of CO$_2$-free water vapour of 1 atm. After two days the temperature of heating was decreased to 900°C for one consecutive day. Then the samples were slowly cooled, crushed and powdered. The
pure end-members CaOHA and SrOHA were also subjected to this heat treatment prior to X-ray diffraction.

Both the Philips Guinier XDC-700 and the Nonius Guinier-de Wolff camera were used, either with CrKα or CuKα. The cell parameters were determined by measuring the position of at least 28 reflections for each sample. The accuracy of the cell parameters a and c of the hexagonal cell which were calculated by using a least squares procedure is estimated to be better than ±0.003 and ±0.002 respectively. The results with both cameras were the same and are represented in Fig. 1. They confirm the linear variation with composition found earlier [3, 4].

In the apatite structure two sublattices occur for the cations. Position I is fourfold (Wyckoff notation f) and position II is six-fold (denoted by h). The structure formula of solid solutions of CaOHA and SrOHA can thus be written as

\[ \text{Ca}_{4+\alpha}\text{Sr}_{\alpha}\text{Ca}_{6-\alpha}\text{Sr}_{1-\alpha} \text{(PO}_4\text{)}\text{b(OH)}_2 \]  

As the cation distributions were not known, we derived them from a comparison of the experimental values of the intensity ratios 210/002, 300/002, 222/002, 002/202, 321/202, 321/410 and 213/312 with their theoretical values depending on α. In the calculation of the theoretical values Lorentz-polarisation and multiplicity factor were taken into account, absorption correction and temperature factor were neglected. Both peak heights and peak areas were taken as a measure for the intensities of reflections.

In this way two sets of seven α values were obtained for each composition x. The average per set was independent of the fact whether peak heights or peak areas were taken. A distribution coefficient K was defined as

\[ K = \frac{[6-(1-\alpha)x][\alpha x]}{[4-\alpha x][1-\alpha x]} \]

and its value and standard deviation were calculated from the two sets of α values at a given x. A constant value for K was found throughout the range 1 ≤ x ≤ 9 within the limits of experimental error. Its overall weighted mean was 0.841 ± 0.065.

This means (a) that apparently an equilibrium is reached according to

\[ \text{Ca(I)} + \text{Sr(II)} \rightleftharpoons \text{Ca(II)} + \text{Sr(I)} \]

at the temperature of preparation and (b) that the solid solutions are close to ideal in the thermodynamical sense [9]. The latter conclusion enables the quantitative estimation of the effect of Sr²⁺ ion incorporation on the solubility product of CaOHA from knowledge about the solubility products of both CaOHA and SrOHA [7].

N.B. The individual values of α and K derived from the intensity ratios are available on request.

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REFERENCES

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Synthesis and characterization of complexes of palladium(II)
and platinum(II) with a tetraaza macrocycle

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Transition metal complexes with synthetic tetraaza macrocyclic ligands have been extensively investigated in recent years. Most of these studies have involved first row transition metals with emphasis on examining the role of the macrocyclic ligand on the