Hexanuclear Iridium Clusters Supported on Magnesium Oxide

S. D. Maloney,1 M. J. Kelley,1 D. C. Koningsberger,§ and B. C. Gates*1

Center for Catalytic Science and Technology, Department of Chemical Engineering, University of Delaware, Newark, Delaware 19716, Engineering Technology Laboratory, Experimental Station, E. I. du Pont de Nemours and Company, Wilmington, Delaware 19898, and Laboratory for Inorganic Chemistry and Catalysis, Eindhoven University of Technology, P.O. Box 513, 5600 MB Eindhoven, The Netherlands

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Introduction

The crucial structural elements in numerous modern materials are microstructures dispersed on surfaces of supports (substrates); the microstructures range from metals in supported catalysts to semiconductors in electronic devices. Understanding of the chemistry and function of these materials will be advanced by the development of methods for synthesis and characterization of discrete, structurally uniform microstructures. Organometallic clusters have been used frequently in attempts to prepare robust metal cluster carbonyls inferred to have given extended X-ray absorption fine structure (EXAFS) spectroscopies, with the data indicating an average Ir-Ir coordination number of 4.1 and an average Ir-Ir distance of 2.77 Å, in close agreement with the published crystallographic data for this anion in the solid state. The reactions of the iridium carbonyls on the basic MgO surface closely parallel those of iridium carbonyl anions in basic solutions. Treatment of the MgO-supported [Ir4(CO)12]- in He followed by H2 at 300°C gives decarbonylated surface species shown by EXAFS spectroscopy to have iridium cluster frames substantially unchanged from those of the carbonyl anion precursor. The decarbonylated iridium clusters are among the most nearly uniform supported metals.

Results

Formation of [Ir4(CO)12]3-/MgO and Characterization by Extraction and Infrared Spectroscopy. When [Ir4(CO)12] was brought in contact with partially hydroxylated MgO, a chemisorbed iridium carbonyl was formed, inferred on the basis of infrared and EXAFS spectroscopies to be [HIr4(CO)11]-; this anion could be removed from the surface by extraction with [PPN][Cl] in THF.4 Treatment of the sample formulated as [HIr4(CO)11]/MgO with CO at temperatures greater than 100°C led to changes in the infrared spectrum (Figure 1). The spectrum characteristic of the supported tetrairidium carbonyl anion disappeared as the temperature was raised to 200°C, and an infrared spectrum (Figure 1C) similar to that of [Ir6(CO)15]2- was observed by Stevens et al.5 as a product of the reaction of [Ir4(CO)12] with Na in THF.6 To be more confident of the spectrum of [Ir6(CO)15]2- in solution, we repeated its preparation from [Ir4(CO)12] with Na in THF by the method of Stevens et al.7 In agreement with the observations of Stevens, the reaction proceeded with the spectrum characteristic of [Ir6(CO)15]2- and a product of its reaction in the presence of Na, [Ir(CO)4]2-, as shown by a band at 1895 cm⁻¹, indicative of the former, and a strong band at 1895 cm⁻¹, indicative of the latter. In contrast to the published spectrum of Stevens et al.,8 another band, at 1991 cm⁻¹, was observed to grow in proportion to the 1895 cm⁻¹ band. However, a splitting into two bands, at about 1995 and 1985 cm⁻¹, was observed by Stevens et al.8b as the [Ir6(CO)15]2- was formed in the presence of Na, in agreement with the results of this work. These two strong bands are inferred to indicate the hexairidium cluster anion. The data of Table I indicate that the cation exerts a significant influence on the spectrum, and the data for the Na+ and PPN+ salts indicate that the presence of the 1991 cm⁻¹ band

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6 The formation of the hexairidium cluster anion was indicated by the appearance of a peak at 180 cm⁻¹, reported by Angoletta et al.6 However, these authors reported the strong band at 1970 rather than 1890 cm⁻¹ for [NEt₄][Ir6(CO)15]. The hexairidium cluster anion continued to react with the sodium to form [Ir(CO)4]2-, which was identified by a characteristic strong band at 1895 cm⁻¹. Stevens et al.1 reported an improved synthesis of the anionic hexairidium cluster, whereby [Ir4(CO)12] reacts with Na in THF under CO until the infrared spectrum of the solution indicates only two species, [Ir6(CO)15]2- and [Ir(CO)4]2-. These authors did not report the locations of all the bands in the infrared spectra attributed to these two anions, but they did attribute a band at 1938 cm⁻¹ to [Ir6(CO)15]2-.
Hexanuclear Ir Clusters Supported on MgO


TABLE I: Infrared Spectra in the Carbonyl Stretching Region of MgO-Supported Species Formed from [Ir₆(CO)₁₅] and Related Molecular Clusters

<table>
<thead>
<tr>
<th>Sample Description</th>
<th>λₜₜ, cm⁻¹</th>
<th>ref</th>
</tr>
</thead>
<tbody>
<tr>
<td>MgO-supported sample prepared from [Ir₆(CO)₁₅]</td>
<td>2056 (w), 2008 (s), 1831 (w)</td>
<td>this work</td>
</tr>
<tr>
<td>[PPN][Cl] extract from above sample</td>
<td>1987 (s), 1979 (s), 1771 (m), 1716 (w)</td>
<td>this work</td>
</tr>
<tr>
<td>[PPN][Ir₆(CO)₁₅] prepared by method of Stevens et al.</td>
<td>2068 (w), 2017 (m), 1993 (s), 1984 (s), 1928 (vw), 1823 (vw), 1788 (w), 1735 (w)</td>
<td>this work</td>
</tr>
<tr>
<td>[Na₂][Ir₆(CO)₁₅]</td>
<td>1993 (s), 1984 (s), 1928 (w), 1788 (m), 1735 (m)</td>
<td>this work</td>
</tr>
<tr>
<td>[NMMe₄][Ir₆(CO)₁₅]</td>
<td>2030 (sh), 1970 (s), 1775 (s), 1735 (s)</td>
<td>this work</td>
</tr>
</tbody>
</table>

Figure 1. Infrared spectra of [Ir₆(CO)₁₅] and the sample prepared by adsorption of [Ir₆(CO)₁₅] on MgO followed by treatment in CO. (A) DRIFTS of sample prepared by adsorption of [Ir₆(CO)₁₅] on MgO followed by treatment with CO at 1 atm and 200 °C for 1/2 h; (B) DRIFTS of sample treated in CO at 15 atm and 200 °C for 2 h in a flow reactor; (C) spectrum of the extract solution formed from bringing the sample of spectrum B in contact with [PPN][Cl] in THF; (D) spectrum of [Na₂][Ir₆(CO)₁₅] in THF solution.

Among the five spectra shown in Figure 1D, the peak at about 1825 cm⁻¹ is typical of metal carbonyls (among others) in the latter may be attributed to strongly ion-paired [Ir₆(CO)₁₅]²⁻ associated with Na⁺ retained from the synthesis of Na and in agreement with the results of Stevens et al. The spectrum also indicates the presence of some residual tetranuclear species, [Ir₄(CO)₁₂] and [Ir₄(CO)₁₅], as shown by small peaks at 2068 and 2017 cm⁻¹, respectively. This spectrum is similar to that characteristic of the [PPN][Cl] solution prepared from the MgO-supported iridium carbonyl extract that had been treated with CO at 200 °C. The spectrum of the Na⁺ salt of [Ir₆(CO)₁₅]²⁻ shown in Figure 1D has a peak at about 1825 cm⁻¹; this peak was not present in every spectrum characterizing this sample, and it is attributed to an unidentified impurity. In summary, the present results are judged to be consistent with the reports of Stevens et al. and Angoletta et al., and the differences in the several spectra that are attributed to [Ir₆(CO)₁₅]²⁻ reflect different ion-pairing effects.

The spectrum of the surface-bound iridium carbonyl species (Figure 1) is similar to that of [Ir₆(CO)₁₅]²⁻ in solution but is shifted and broadened with respect to the solution spectrum, as is typical of metal carbonyls on solid supports. A peak attributed to unconverted tetranuclear iridium is also present, as in the solution spectrum (Table 1). All these results are consistent with the inference that [Ir₆(CO)₁₅]²⁻ had formed on the MgO surface and was the predominant organometallic species, but they are not sufficient to define the surface structures. Extraction with [PPN][Cl] evidently took place by simple cation metathesis, and the results of the extraction experiment are indicative of [Ir₆(CO)₁₅]²⁻ as the predominant surface species.

Characterization of [Ir₆(CO)₁₅]²⁻/MgO by EXAFS Spectroscopy. EXAFS spectra of the supported iridium carbonyl were measured to test the structural inferences based on infrared spectroscopy. The raw EXAFS data characterizing the sample prepared by adsorption of [Ir₆(CO)₁₅] on MgO followed by treatment with CO at 200 °C and inferred from the infrared spectra to incorporate [Ir₆(CO)₁₅]²⁻ are shown in Figure 2B. For comparison, the raw EXAFS data characterizing the sample prepared by adsorption of [Ir₆(CO)₁₅] on MgO to give the stable [Ir₄(CO)₁₂]⁺/MgO are shown in Figure 2A. In the intermediate (8 < k < 11 Å⁻¹) and higher ranges of the wave vector k, there are strong oscillations characteristic of the sample inferred to incorporate the hexaergic clusters. The intense oscillations at the higher values of k are characteristic of metal–metal interactions, consistent with the presence of clusters, and the high frequency of these in comparison with those of [Ir₄(CO)₁₂]⁺ indicates the presence of Ir–Ir bonds with greater bond distances than in [Ir₄(CO)₁₂]⁺.

The Koningsberger difference file technique was used to analyze the EXAFS data obtained for the sample treated with flowing CO at 1 atm and 200 °C for 8 h. The raw data were Fourier transformed with k² weighting in the range 2.77 < k < 15.03 Å⁻¹. The first-shell contributions to the EXAFS were isolated by inverse Fourier transformation of the data over the range 1.08 < R < 3.04 Å, where R is distance from the scattering atom (Ir). In agreement with the infrared spectrum, the EXAFS data show that there was predominantly one symmetrical metal carbonyl species present.

The isolated first-shell EXAFS data were fitted on the basis of reference files determined experimentally from EXAFS data for Pt foil (to represent the Ir–Ir interactions) and [Ir₄(CO)₁₂] (to represent the Ir–CO interactions). The procedure was started by fitting the isolated first-shell data over the higher k range on the basis of the reference file for the Ir–Ir interactions. After this fit had been determined, an EXAFS function (χ function) was calculated from the parameters that had been estimated and subtracted from the isolated first-shell data. If the estimate were a good first approximation, the residual function would consist predominantly of the contributions arising from the Ir–C and Ir–O interactions, where O refers to the carbonyl oxygen. The next step was to fit the residual function with contributions on the basis of the reference file characterizing these interactions (the [Ir₄(CO)₁₂] file). Fitting the residual function with the two contributions (Ir–C and Ir–O) was not sufficient to represent the data in the region of approximately 2.1 Å. Therefore, in subsequent fitting, another contribution was included, representing a carbon atom of a bridging carbonyl ligand (Ir–Cₙ).

After estimation of the coordination parameters for the Ir–Ir and Ir–CO interactions, further fine-tuning of the fit was carried out with contributions, Ir–Cₙ (t refers to the terminal carbon), Ir–Cₚ, Ir–Oₚ, and Ir–Ir, for the first-shell data. The parameters for one contribution were allowed to vary in the fitting routine while parameters for the other three contributions were held constant. Alternatively, the coordination numbers for all the contributions were allowed to vary while the remaining parameters were fixed. This iterative procedure was followed, in turn, by varying all the other parameters, namely, the Deby–Waller factors, distances, and inner potential corrections. The fitting

procedure was performed repeatedly until further optimization resulted in minimal changes in the sum of the squares of the residuals between the isolated first-shell data and the calculated fit. The parameters determined in the final fit are shown in Table II; estimates of the experimental uncertainties in the parameters are shown in the table.

The results indicate an Ir-Ir coordination number of approximately 4, consistent with the earlier inference that the predominant surface species was a hexairidium carbonyl cluster. The average Ir-Ir distance is 2.77 Å.

There is a contribution from a terminal carbonyl carbon atom with a coordination number of approximately 2.5 at a distance of 1.85 Å. A contribution inferred to be a bridging carbon is indicated at a distance of 2.13 Å with a coordination number of 1.5. A fourth contribution is identified as a terminal carbonyl oxygen present at a distance of 3.05 Å with a coordination number of about 1.5.

To verify that the fit was approximately equally good with respect to the low-Z (C and O*) and high-Z (metal) scatterers, the final fit was compared with the first-shell data with both \( k^2 \) and \( k^4 \) weighting (Figure 3). The \( k^4 \)-weighted plot emphasizes the contributions of the low-Z neighbors, and the \( k^2 \)-weighted plot is dominated by the metal-metal interactions. The fit was checked in the Fourier transform as well as the \( \chi \) plot. The fits of the individual contributions are shown in the phase-corrected Fourier transforms along with the Fourier transform of the residual \( \chi \) plot after subtraction of all the other contributions (Figure 3). For example, to verify that the Ir-C\textsubscript{t} contribution was fitted properly, the Fourier transform of the calculated EXAFS for this contribution was compared with the Fourier transform of the residual function obtained by subtracting the Ir-Ir, Ir-C\textsubscript{p}, and Ir-O* contributions from the total. This comparison shows that the fit for the whole spectrum is a sum of contributions which are individually fitted well.

The number of adjustable coordination parameters which can be used in the fitting procedure to give statistically significant results is given approximately by the Nyquist theorem as \( 2\Delta k\Delta R/\pi \), where \( \Delta k \) and \( \Delta R \) are the ranges of the forward and inverse Fourier transformations. For the data analysis presented here, this number is calculated to be 15.3 (\( \Delta k = 12.3 \) Å\(^{-1} \) and \( \Delta R = 1.96 \) Å), which is close to the total number of parameters (16) used in the fit. The goodness of the fit is shown by the comparisons of Figure 3.

Decarbonylation of \([\text{Ir}_6(\text{CO})_{12}]^{2-}\)/MgO. Qualitative temperature-programmed decomposition experiments indicated that the surface-bound iridium carbonyl is decarbonylated at approximately 300 °C. Consequently, the sample treated in He at this temperature was characterized by infrared spectroscopy; after 2 h, only a very weak and broad carbonyl absorption remained, indicating that virtually all the CO ligands had been removed.

Since this decarbonylation could be expected to form hydrocarbon products in addition to CO and possibly CO\(_2\), the sample was treated in H\(_2\) at 1 atm and 300 °C for 2 h in an attempt to remove any hydrocarbons and leave the clean decarbonylated clusters on the support in the form of iridium hydrides. Since \([\text{Ir}_6(\text{CO})_{12}]^{2-}\) has a relatively stable metal frame, the goal of these treatments was to decarbonylate the cluster while leaving the metal frame intact and giving a structurally simple supported metal, Ir\(_6\). To test whether this goal was achieved, the decarbonylated sample in the presence of H\(_2\) at approximately liquid-nitrogen temperature, was characterized by EXAFS spectroscopy. The full EXAFS analysis, carried out by essentially the same method

\[ \text{Ir-Ir} \]

\[ 4.08 \pm 0.02 \]
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The results are consistent with the inference that the iridium carbonyl species initially formed from \([\text{Ir}_4(\text{CO})_{12}]\) on the MgO support were converted in high yield to \([\text{Ir}_4(\text{CO})_{13}]^-\). The evidence includes (1) the infrared spectra of the supported iridium carbonyls, which clearly demonstrate the conversion of the initially formed species but do not give an unequivocal indication of the product, (2) the infrared spectrum of the species extracted from the solid with [PPN][Cl], which is consistent with the identification of \([\text{Ir}_4(\text{CO})_{13}]^-\) as the predominant extracted species (which, however, was present with other iridium carbonyl anions and which reacts in the extract solution to give \([\text{Ir}_4(\text{CO})_{13}]^-\)) and suggests that this cluster was the predominant organometallic species on the surface, and (3) the EXAFS spectrum of the surface-bound iridium carbonyl.

The EXAFS results provide the strongest evidence; they show that, in the sample prepared from \([\text{Ir}_4(\text{CO})_{12}]\) on MgO after treatment with CO at 200 °C, the Ir neighbors included Ir and low-Z scatterers identified as components of carbonyl ligands. The relative contribution of each of these neighboring atoms must be explained to support the identification of the surface species as \([\text{Ir}_6(\text{CO})_{15}]^-\).

All the Ir atoms in crystalline \([\text{Ir}_6(\text{CO})_{15}]^-\) are stereochemically equivalent, with each bonded to two terminal carbonyl ligands and one bridging carbonyl ligand in addition to four Ir atoms. The average Ir-Ir distance is substantially longer (2.773 Å) than those in \([\text{Ir}_4(\text{CO})_{13}]\) (2.69 Å) and \([\text{Ir}_4(\text{CO})_{12}]^-\) (2.73 Å).

The EXAFS results for the supported iridium Carbonyl match the crystallographic data for \([\text{Ir}_6(\text{CO})_{15}]^-\) within experimental error: the EXAFS data indicate four Ir neighbors at an average distance of 2.77 Å; the agreement is striking.

The EXAFS results are not sufficient to determine fully the carbonyl ligation in the supported cluster, but they can be interpreted satisfactorily on the basis of multiple scattering effects. The difference between the coordination numbers corresponding to the Ir-C, (2.5) and Ir-O* (1.3) interactions can be understood on this basis, as follows: The positions of the oxygens of the bridging and terminal carbonyl ligands in the crystalline cluster are known. The distances between the metal and oxygen atoms of the bridging and terminal carbonyl ligands are equivalent; however, the Ir-C-O angle is 176.2°, whereas the Ir-C-Oa angle is 137.8°. When the bond angle is greater than approximately 140°, the EXAFS phase shift and amplitude function depend strongly on the positions of the atoms. The terminal Ir-O* interactions in \([\text{Ir}_4(\text{CO})_{13}]^-\) [Os$_3$(CO)$_{13}$]$_2$ and [H$_2$Re$_2$(CO)$_{15}$]$_2$ illustrate the effect for metal-C-O bond angles close to 180°. In contrast, the multiple scattering effect is not significant for the bridging carbonyl oxygen in crystalline \([\text{Ir}_4(\text{CO})_{12}]^-\), since the Ir-C-O angle is only 137.8°.

Since the terminal and bridging carbonyl oxygen atoms are located at the same distance from Ir, it is difficult to separate the two contributions and to determine the overall Ir-O* coordination number. Consequently, we infer that the discrepancy between the total coordination number determined by the EXAFS data for the Ir-O* contribution (1.3) and the expected coordination number (3) is a result of there being two types of carbonyl oxygen contributions at the same distance but experiencing different degrees of multiple scattering that lead to different phase shifts, which may give a negative interference and a low estimate of the overall coordination number.

Figure 3. Imaginary parts of the Fourier transforms of the individual contributions to the EXAFS characterizing the sample prepared by adsorption of \([\text{Ir}_4(\text{CO})_{12}]\) on MgO followed by treatment with CO at 1 atm and 200 °C for 2 h. (A) Residual Ir-Ir contribution (---) and best fit (--), Ir-C contribution (---) and best fit (--); (B) residual Ir-Ir contribution (-) and best fit (--), k$^2$-weighted and Pt-Pt phase corrected; (C) residual Ir-Ir contribution (---) and best fit (--), k$^2$-weighted and Ir-C phase corrected; (D) residual Ir-O contribution (---) and best fit (--), k$^2$-weighted and Ir-O* phase corrected.

Described above, is complicated by the need to represent the metal-support interface to give a good fit; details will be reported elsewhere. The important result is that the cluster carbonyl ligands had been removed without much change in the cluster nuclearity, presumably as a cluster hydride was formed; the Ir-Ir near-neighbor distance is close to that of bulk Ir.\(^{13}\)

\(^{11}\) van Zon, F. B. M.; Maloney, S. D.; Gates, B. C.; Koningsberger, D. C. To be published.
The formation of this cluster anion is one of the few examples of anion reactions of metal carbonyl anions. Determination of reactions of metal carbonyl anions can be used to determine the carbonyl ligand environment, but the data are not sufficient to determine the structure, but they are consistent with those of the crystalline cluster when interference on the basic surface closely parallels the chemistry of iridium carbonyls in basic solution.\(^{(19)}\)

We conclude, therefore, that the predominant surface species is likely a mixture of iridium hydrides, is regarded as a nearly intact during the decarbonylation. The resulting surface species, represented here as taking place on the basic surface closely parallels the chemistry of iridium carbonyls in basic solution.\(^{(19)}\)

The hexairidium cluster anion is robust, and when the MgO-supported indium carbonyl was prepared by slurring [Ir\(_4\)(CO)\(_{12}\)] (0.029 g) in freshly distilled hexanes (50 mL) in a Schlenk flask under a N\(_2\) blanket. This calcined MgO (2.0 g) was added to the slurry. The mixture was stirred under N\(_2\) for 4 h at room temperature then dried by evacuation for 14 h. The transformations of the iridium carbonyls on the MgO surface are summarized in Scheme I; the “molecular” chemistry as described above was loaded, as well as several pieces of Na were added. The mixture was stirred for several hours, and the solution spectrum monitored by infrared spectroscopy. Infrared spectroscopy. Transmission infrared spectra of liquids were collected with a Nicolet 7199 spectrometer. The spectra of powders were collected with the instrument operated in the DRIFTS (diffuse reflection infrared Fourier transform spectroscopy) mode; the cell, part of a flow system, was equipped with a sapphire window. Sample handling was done in the drybox.

The infrared experiments with the solid were performed after connection of the cell to a gas manifold allowing flow of He, N\(_2\), CO, and H\(_2\). After removal of the residual air from the flow system, the flow to the cell was begun; flow rates were 20–30 mL/min. The sample was typically heated to a set temperature, allowed to reach steady state in 1/2 h, and then scanned 32 or more times; the signal was averaged. The sample scans were ratioed to the signal from dried KBr (512 scans) and the Kubelka–Munk function calculated.

**Experimental Methods**

**Sample Preparation.** Samples were prepared under N\(_2\) in a vacuum atmosphere glovebox and on a Schlenk vacuum line. [Ir\(_4\)(CO)\(_{12}\)] (Strem) was used without further purification. Reagent grade hexanes were dried by refluxing over sodium benzophenone ketyl. N\(_2\) with a purity of 99.999% was purchased from Matheson and passed through beds of CuO and molecular sieve to remove traces of O\(_2\) and moisture. CO was purchased from Matheson (UHP grade) and passed through a bed of activated carbon heated to a temperature >200 \(^\circ\)C to remove traces of iron carbonyl contaminants and through a bed of molecular sieve particles to remove moisture.

Partially dehydroxylated MgO powder was prepared as the powder MgO (MCB) in flowing O\(_2\) was heated to 400 \(^\circ\)C and then held at that temperature for 2 h. The MgO was then placed under vacuum (ca. 10\(^{-3}\) Torr), held at 400 \(^\circ\)C for 14 h, and cooled under vacuum to room temperature and unloaded in the glovebox. The sample had a surface area of 75 m\(^2\)/g, determined by N\(_2\) adsorption.

The MgO-supported iridium carbonyl was prepared by slurring [Ir\(_4\)(CO)\(_{12}\)] (0.029 g) in freshly distilled hexanes (50 mL) in a Schlenk flask under a N\(_2\) blanket. This calcined MgO (2.0 g) was added to the slurry. The mixture was stirred under N\(_2\) for 4 h at room temperature then dried by evacuation for 14 h at room temperature.

[Ir\(_4\)(CO)\(_{12}\)] was prepared by the method of Stevens et al.\(^{(19)}\) as [Ir\(_4\)(CO)\(_{12}\)] was slurried with THF under CO, and several pieces of Na were added. The mixture was stirred for several hours, and the solution spectrum monitored by infrared spectroscopy.

**Infrared Spectroscopy.** Transmission infrared spectra of liquids were collected with a Nicolet 7199 spectrometer. The spectra of powders were collected with the instrument operated in the DRIFTS (diffuse reflection infrared Fourier transform spectroscopy) mode; the cell, part of a flow system, was equipped with a sapphire window. Sample handling was done in the drybox.

The infrared experiments with the solid were performed after connection of the cell to a gas manifold allowing flow of He, N\(_2\), CO, and H\(_2\). After removal of the residual air from the flow system, the flow to the cell was begun; flow rates were 20–30 mL/min. The sample was typically heated to a set temperature, allowed to reach steady state in 1/2 h, and then scanned 32 or more times; the signal was averaged. The sample scans were ratioed to the signal from dried KBr (512 scans) and the Kubelka–Munk function calculated.

**X-ray Absorption Spectroscopy.** Sample Preparation and Data Collection. The samples, some having undergone treatments in gas atmospheres, were pressed into wafers. The MgO-supported sample prepared from [Ir\(_4\)(CO)\(_{12}\)] as described above was loaded, under N\(_2\), into a tubular stainless-steel reactor and treated under flowing CO at 1 atm. With a flow rate of about 30 mL/min, the sample was heated to 200 \(^\circ\)C at a rate of 5 \(^\circ\)C/min and then held at 200 \(^\circ\)C for 8 h; the sample was cooled to room temperature under CO flow. At room temperature, the reactor was flushed with He for 3 min, and the sample was unloaded into a vial inside the drybox. The sample was placed inside three layers of glass vials, each individually sealed with paraffin, the vial cap, and electrical tape. All these steps were carried out at the University of Delaware.

The packaging precautions were taken to allow transport of the sample, without air contamination, to the Synchrotron Light...
Cation Location in La,Na–Y Zeolites by Two-Dimensional 23Na Nutation NMR

Ching-Fang Lin and Kuei-Jung Chao

Department of Chemistry, Tsinghua University, Hsinchu, Taiwan, Republic of China
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The variation of the cation distribution with the lanthanum-exchanged level in hydrated La,Na–Y zeolites was clearly monitored by two dimensional nutation sodium-23 NMR, in which the chemical shift and the quadrupole interaction could be separated. From a comparison of 23Na NMR at different hydrated states and temperatures it is concluded that the mobility of water molecules and sodium ions can be reduced by decreasing sample temperature and the migration of lanthanum ions from supercages to small cages causes the redistribution of sodium ions after dehydration at 350°C.

Acknowledgment. We thank Professor W. L. Gladfelter for helpful comments about the infrared spectra and ion pairing and for providing unpublished infrared spectra. This research was supported by the National Science Foundation (CTS-8910633 and CTS-9012910) in the U.S.A., by the NWO in the Netherlands, and by a NATO travel grant.


TABLE IV: Structural Parameters for the Reference Compounds and the EXAFS Parameters for the Preparation of the Reference Files

<table>
<thead>
<tr>
<th>Reference compd</th>
<th>Shell</th>
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<td>Pt foil</td>
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<td>Ir₄(CO)₁₂</td>
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<td>Ir-O*</td>
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<td>2.00-16.45</td>
<td>1.98-3.30</td>
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</table>

*Notation: n, power of k used in weighting of Fourier transform; k range, limits of forward Fourier transformation; R range, limits of shell isolation [Ir-Ir contribution (N = 3, R = 2.69 Å) previously subtracted].

The EXAFS data were collected at X-ray beamline 9.2 with Pt foil and [Ir₄(CO)₁₂] were used as references for analysis of the EXAFS data. The reference data were collected at beamline 9.2 at Daresbury. The Ir-Ir contributions were characterized with an experimentally determined reference file obtained from EXAFS data for the Pt foil. The interactions between the Ir atoms and the terminal and bridging carbon atoms, Ir-C₁ and Ir-C₄, respectively, were analyzed with the reference file characterizing the Ir-C interaction in [Ir₄(CO)₁₂]. The interactions between Ir and O* were also analyzed with a reference file obtained with [Ir₄(CO)₁₂]. The details of the generation of these reference files are described elsewhere.22 The structural parameters of the reference materials and the EXAFS parameters reported for the preparation of the reference files are shown in Table IV.

Introduction

The acidity and thermal stability of synthetic Na–Y zeolite can be dramatically improved by ion exchange with lanthanum ions, and the amount of improvement is controlled by the degree of exchange and therefore the location of the cations.1,2 The structural characteristics and the catalytic application of Na–Y zeolites have been extensively investigated; most of the published structural studies were done on a lanthanum-exchanged Y zeolite under different heat treatments.3–5 Few studies were found on the effect of the degree of cation exchange on the cation distribution in Y zeolites.10–12 The determination of the cation distribution has involved a number of techniques, e.g., X-ray and neutron diffractions,3–5 infrared,6,7 and NMR.8,13 With use of diffraction methods, the stationary cations can be located, and about half of cations can often be located for hydrated samples by X-ray diffraction. With the IR method, the vibrational modes of framework and the adsorbed molecules are usually investigated; however, only limited information was obtained for the distribution of cations in zeolites. Our previous 29Si NMR study8,10 showed that the variation of the cation location and occupancy at different sites led to the change of the electric field on framework silicon atoms around cation sites and hence the change of corresponding interactions.