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High-pressure Raman investigation of mutual solubility and compound formation in Xe-N₂ and Ne-N₂

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In this study, the mixed solid phases of the molecular systems Xe-N₂ and Ne-N₂ have been investigated up to 13 GPa using Raman spectroscopy and so-called ν-T and p-T scans. Both systems show a rich variety of solid phases. In Xe-N₂ a van der Waals compound has been found. The solubility of N₂ in solid Xe is about 7.5 mol%, while the solubility of the larger Xe atoms into the β and δ phases of N₂ is higher. In Ne-N₂ two van der Waals compounds have been detected. The Raman spectrum of one of these compounds is very similar to that of He(N₂)₁₁. N₂ does not dissolve into solid Ne, while Ne does dissolve into the β and δ phases of N₂, even though the atomic diameter of Ne is considerably smaller than the molecular one of N₂. Analogous to Ar-N₂, the Ne and Xe atoms are mainly located at the α sites in the δ* phase. For Ne-N₂, in contrast to Xe-N₂, the β*-δ* transition shifts to higher pressure. The ε*–δ* transition is not shifted for Ne-N₂. The results have been compared with computer simulations and analytical theories for hard-sphere systems. It is shown that the ratio of the diameters of the molecules, α, is only a very rough indication for the phase behavior of mixed molecular systems at high pressure. [S0163-1829(99)08641-5]

INTRODUCTION

In molecular systems the introduction of a second component leads to interesting phenomena like the formation of van der Waals compounds, orientational glasses, and disordered mixed solids. The properties of these new solids are quite different from those of the constituent components. In addition, phase lines can shift enormously, so that certain high-pressure phases can be formed at considerable lower pressures in the mixed system, and other phases might not be formed at all.

The literature about the mutual solubility of simple molecular systems in the solid phase is rather limited. More work has been done on metallic systems. It is generally believed that solubility at high density, in particular in solid systems, is mainly governed by geometrical effects, as expressed by the well-known Hume-Rothery rule.¹ This empirical rule states that a binary mixed solid is obtained only if the ratio of the diameters of the molecules, α, is larger than 0.85. As long as the mixed solid is completely disordered, the Hume-Rothery rule is in fair agreement with computer simulations² and analytical theories.³,⁴ In addition, the simulations show that the solubility of large spheres in the solid formed by small spheres is much smaller than vice versa. We will investigate whether these geometrical rules are also valid for molecular systems.

When the molecular diameters of the two components differ considerably, stoichiometric compounds can be formed. For hard-sphere systems⁵–⁸ these compounds (AB, AB₁₂, and AB₁₁₃) are stabilized by the entropy and efficient packing.

Up to now, a number of compounds have been found in molecular systems. Because the interactions between the molecules are of the van der Waals type, these solids are called van der Waals compounds. In the literature the occurrence of these van der Waals compounds is often explained by efficient packing of hard spheres, dictated mainly by α (see, e.g., Ref. 9). In this work we will investigate if this is indeed the case (see also Ref. 10).

Binary mixtures of nitrogen and a noble gas are suitable model systems for the following reasons. The nitrogen molecule is a simple diatomic slightly nonspherical molecule, the interaction potential is well known, and pure nitrogen has been studied extensively (see, e.g., Ref. 11 and references therein). It provides a good opportunity to investigate the effect of asphericity and weak intermolecular forces on the mutual solubility and compound formation. Moreover, nitrogen has a peculiar high-pressure phase, the δ phase (Pm±3n), in which the molecules on the a sites are completely orientationally disordered (spheres), whereas the molecules on the c sites are orientationally disordered only in a plane normal to the faces of the unit cell (disks). The system shows a second-order transition, δ → δₕ.¹²

Previous investigations of He-N₂ and Ar-N₂ revealed a very complex phase behavior for these mixtures. In He-N₂ (α=0.62), a disordered solid solution exists at high pressure and low temperature on the nitrogen-rich side of the phase diagram.¹³,¹⁴ A stoichiometric compound [He(N₂)₁₁] has been detected at room temperature.¹⁵ The Raman spectrum of the compound is very similar to that of ε-N₂,¹⁶,¹⁷ but the intensity ratio between the two Raman peaks ν₁ and ν₂ is decreased compared to the spectrum of pure ε-N₂.

The high-pressure investigations on the mixture Ar-N₂ (Refs. 18 and 19) (α=0.95, with nitrogen the larger molecule) show that the solubility of argon in the solid β and δ phases of nitrogen is very high. The Raman spectra reveal that the argon atoms have a preference for the a sites in the δ phase. The distribution of the argon atoms over the sites is temperature and pressure dependent. The δ-ε transition has not been found experimentally in the mixed system. As shown by computer simulations, it probably shifts towards lower temperatures or higher pressure by the addition of argon.²⁰
In order to study the evolution of the phase diagram, the mutual solubility, and the compound formation as a function of $a$ in nitrogen systems, we have studied the phase diagrams of Ne-N$_2$ ($a=0.74$) and Xe-N$_2$ ($a=0.89$).

We will present Raman measurements as a function of pressure up to 13 GPa. In addition, so-called $\nu$-$T$ and $p$-$T$ scans will be performed for Ne-N$_2$. Finally, phase diagrams for these mixtures will be proposed and the results will be discussed.

**EXPERIMENTAL RESULTS**

Xe-N$_2$ mixtures, with nitrogen mole fractions $x$ of 0.075 and 0.84 and Ne-N$_2$ mixtures with $x=0.90$, 0.35, and 0.052 have been prepared in a gas compressor. The experimental procedure has been described elsewhere. The pressure is determined by the ruby luminescence method using the pressure scale of Ref. 21 and the temperature correction of Ref. 22. The Raman spectra have been fitted by Lorentzian curves. These curves are used to determine the peak position and the full width at half maximum (FWHM). The instrumental width is about 0.75 cm$^{-1}$.

In all cases, only one single Raman peak ($\nu_F$) has been observed at low pressures in the fluid phase, as could be expected. Observations through the microscope show that the sample space is completely transparent under these conditions. In addition, visual observations show the appearance of a solid phase when the pressure is increased.

As is shown by visual observations, usually we do not obtain a fine-grained phase-separated mixture, but larger pieces (compared to the focal cylinder of the laser) of the individual solid phases are formed. Therefore, when the laser is focused at different spots in the sample space and the intensity ratio between Raman peaks does not change, it is assumed that these peaks are all modes of one single solid phase. Important information is also obtained from the line-width, since it is not only a function of pressure, but it also depends strongly on the composition. Moreover, the line-width is much smaller in a solid than in a liquid at the same pressure and temperature.

**Xe-N$_2$**

The measurements have been performed at constant temperature (408 K). Typical Raman spectra are shown in Figs. 1 and 2. The vibrational frequency as a function of pressure is presented in Fig. 3.

For $x=0.84$, at the fluid solid transition (4.9 GPa), a drastic drop of the line-width occurs similar to that observed for pure N$_2$. For $p\geq 4.9$ GPa, the Raman spectrum consists of two peaks as shown in the lower spectrum of Fig. 1. For reasons to be explained later, the high-frequency peak is referred to as $\nu_{bs}$ and the low-frequency peak as $\nu_F$. At 6.0 GPa three additional vibrational modes were measured (Fig. 1) indicating a three-phase equilibrium. The additional modes consist of one peak on the high-frequency side of $\nu_{bs}$ ($\nu_{1a,bs}$) and two peaks on the low-frequency side of $\nu_{bs}$ ($\nu_{2a,bs}$ and $\nu_{2b,bs}$). Above 6.0 GPa, $\nu_{bs}$ and $\nu_F$ could no longer be detected, but an additional peak ($\nu_{1b,bs}$) below $\nu_{1a,bs}$ was found. The four peaks are all modes of one single phase ($\delta^b$). The intensity of the peaks $\nu_{1a,bs}$, $\nu_{1b,bs}$, and $\nu_{2b,bs}$ divided by that of $\nu_{2a,bs}$ is, respectively, 0.25±0.02, 0.03±0.01, and 0.06±0.03. The mode $\nu_{1b,bs}$ is no longer detectable above 9.8 GPa.

For $x=0.075$ coexisting solid and fluid phases have been observed above 0.83 GPa. The amount of solid increases at increasing pressure, until the sample space seems to be completely filled with solid. A typical Raman spectrum of the fluid phase is represented by the lower spectrum in Fig. 2. Note that the small peak at about 2330 cm$^{-1}$ ($\nu_0$) is the Raman peak of N$_2$ in the ambient air. Above the fluid-solid transition two vibrons are observed as shown at 1.4 GPa in Fig. 2. One can clearly observe that the distance between these two peaks increases as a function of pressure due to a rapid increase of $\nu_F$. Above 1.6 GPa the high-frequency...
peak is no longer detectable. At 5.3 GPa a new peak appears ($n_{S1}$), with a frequency about 12 cm$^{-1}$ higher than $n_{Xe^*}$.

One can observe in Fig. 3 that $n_{S1}$ increases slightly more as a function of pressure than $n_{Xe^*}$.

**Ne-N$_2$**

1. **Constant temperature measurements**

The measurements have been performed at room temperature (296 K), except for one experimental run which was performed at 408 K. Typical Raman spectra are given in Fig. 4. The frequency and linewidth as a function of pressure are presented in Figs. 5 and 6, respectively.

For $x=0.052$, 0.35, and 0.90, the solid-fluid coexistence region is entered, respectively, at 5, 7.71, and 4.44 GPa at room temperature. For $x=0.052$ and 0.35, in the solid-fluid coexistence region, only one Raman peak is observed as shown at 8 GPa in Fig. 4. There is no discontinuity in the peak position or in the FWHM of the Raman spectrum at the entrance point to the solid-fluid region. In the solid-fluid coexistence region the values of the peak position and of the FWHM, and their pressure dependence, are the same for the mixtures $x=0.052$ and 0.35. Therefore, the Raman peak in the coexistence region is attributed to the fluid phase. Thus nitrogen does not dissolve in solid neon.

For the mixture with $x=0.052$, these experiments were also performed at 408 K; essentially the same behavior was observed. At this temperature, the solid-fluid coexistence region is entered at 8.3 GPa.

For $x=0.90$, the Raman spectrum consists of two peaks ($n_F$ and $n_{B^+}$) in the solid-fluid coexistence region. The Raman frequency of $B^+$ is 0.2 cm$^{-1}$ higher than that of pure $B-N_2$ in the stability range of $B-N_2$. The broad Raman peak $n_F$ is present until at least 7.5 GPa, but we could not find the fluid peak for 5.4 GPa<$p<$7.5 GPa possibly because of the low amount of fluid present in the sample space.

Increasing the pressure, in the mixture with $x=0.052$, in addition to $n_F$, two new modes appear ($n_{S2,a}$ and $n_{S2,b}$) at 8.35 GPa at the high-frequency side of $n_F$. The intensity

![FIG. 3. Raman frequency vs pressure of N$_2$ in Xe-N$_2$ at 408 K. $x=0.84$: crosses, homogeneous fluid; squares, $n_{B^+}$; triangles, modes of $\delta^*$. $x=0.075$: plus signs, homogeneous fluid; open circles, fluid in coexistence region; solid circles, $S_{Xe^*}$; diamonds, $S_1$. Solid lines represent the extrapolated frequencies of the fluid (Ref. 33). $\beta$ and $\delta$ phases of pure N$_2$ at 408 K (Ref. 36). Dotted lines are guides to the eyes.](image1)

![FIG. 4. Raman spectra of N$_2$ in Ne-N$_2$ at 296 K. $x=0.35$.](image2)

![FIG. 5. Raman frequency vs the pressure of N$_2$ in Ne-N$_2$ at 296 K. Circles, $x=0.052$: squares, $x=0.35$: triangles, $x=0.90$, open symbols with dot, fluid in homogeneous region; open symbols without dot, fluid in coexistence region; plus signs ($x=0.35$) and crosses ($x=0.90$), Raman frequencies obtained at 296 K in a heating scan. Solid lines represent the frequencies of the fluid (Ref. 33), $\beta$ and $\delta$ phases of pure N$_2$, respectively (Ref. 36) and dashed lines the extrapolated frequencies of pure $v-N_2$ at 296 K (Ref. 36).](image3)
FIG. 6. FWHM vs the pressure of the Raman peak of N₂ in Ne-N₂ at 296 K. Circles, x = 0.052; squares, x = 0.35; triangles, x = 0.90; open symbols with dot, fluid in homogeneous region; open symbols without dot, fluid in coexistence region; solid up triangles, solid down triangles (x = 0.90); S₂(ν₁) and S₃(ν₁); solid squares (x = 0.35), S₂(ν₃) and S₃(ν₃); plus signs (x = 0.052), modes of S₂: crosses (x = 0.35), modes of S₂. For the sake of clarity the linewidths of the other modes have not been plotted.

In the mixture with x = 0.35, there is no longer a ν₁ peak at 9.2 GPa, but two new peaks have been found (ν₁,ν₂, and ν₁,ν₂), with an intensity ratio of 0.24 ± 0.03, as shown in Fig. 4. Through the microscope two different solid phases can be observed. The other solid, which can be observed by visual observations, is not Raman active and must be pure neon. The Raman frequencies ν₁,ν₂ and ν₁,ν₂ can be measured until at least 9.5 GPa. The intensity ratio remains constant over this pressure range. When increasing the pressure further, a change in the Raman spectrum has been observed at 9.8 GPa. The spectrum now consists of five new peaks. Two of these peaks are the modes of S₂, which were already mentioned before. The other three peaks will be referred to as ν53,a, ν53,b, and ν53,c. The intensity ratio between ν53,a and ν53,b is equal to 0.07 ± 0.01 and that between ν53,a and ν53,b is 0.03 ± 0.01. The frequency of ν53,b is within experimental accuracy the same as the extrapolated frequency ν₁,ν₂. The mode ν₁,ν₂ is no longer detected.

In the mixture with x = 0.90 at 11.0 GPa, an additional peak appears on the high-frequency side of ν₁,ν₂, which is referred to as ν53,a, since the frequency is the same as that observed for x = 0.35. The spectrum consists of three peaks, namely, ν53,a, ν1,ν₂, and ν2,ν₆. One would also expect to measure ν53,c, but the intensity of this peak is possibly too low to be detected. Above 11.6 GPa, ν53,a is no longer detected.

2. ν-T and p-T scans

In order to obtain more information about the character of the various solid phases, we have also performed scans of ν and p against temperature for x = 0.35 and 0.90 in steps of about 5 K. This is a very suitable method to detect three-phase lines in binary systems, since generally at a three-phase line the pressure increases or decreases sharply and/or a drastic change is observed in the Raman spectrum. The formation of the various solid phases in Ne-N₂ shows a large metastability or hysteresis. When, for example, for x = 0.35 the temperature was decreased at 7.8 GPa until 220 K and then increased again, the modes of δ and S₂ were measured at and above 9.8 GPa in the mixture, which were already mentioned before. The other three solid phases can be observed. The other solid, namely, S₃, is 0.03, as shown in Fig. 4. Through the microscope two different phases are present, it is identified as a point on the three-phase line S₁-S₂-δ* (Fig. 7a). The pressure increases slightly at this particular temperature. When further increasing the temperature, first the modes of S₂ disappear, then the two peaks of δ* disappear, and a small broad peak appears, ν₆. This peak is identified as a point on the three-phase line S₂-δ*-F (Fig. 7b). We were unable to determine the frequency ν₆ for all temperatures. At even higher temperatures, the pressure increases sharply as shown in Fig. 7a. For reasons to be explained later, this line is denoted as S₂-F-S₃. Note that the Raman frequencies also increase at this three-phase line [Fig. 7(b)], due to the increase of the pressure. When further increasing the temperature, visual observations show the melting of a solid (S₃). This solid is not Raman active and, therefore, must be pure neon. The temperature is now increased very slowly to determine the temperature of the final solid disappearance (S₃*-F point) accurately by visual observations. In Fig. 7b these points are marked by arrows. No discontinuity of the Raman frequency occurs within experimental accuracy at the S₃*-F point. In addition, the FWHM also does not show a discontinuity at this temperature.

x = 0.90. One heating scan has been performed around 11 GPa. The frequencies have been compared in Fig. 10 with those obtained from a heating scan for x = 0.35 at the same pressure, starting at 278 K. As shown in Fig. 8, a small wing (marked by arrow) on the high-frequency side of ν53,a is visible. Above 196 K the mode ν1,ν₂ is no longer detectable. When the laser is focused at a different spot in the sample space, the intensity of ν53,a increases slightly, while the in-
The intensity of $n_{S_3}$ decreases slightly and vice versa. The intensity ratio of $n_{S_3,a}$ and $n_{S_3,b}$ is much larger than for $x=0.35$. In contrast to the results for $x=0.35$, for $x=0.90$, the modes of $S_2$ have not been detected at low temperature. The two peaks which appear at 310 K ($n_{S_2,a}$ and $n_{S_2,b}$) have a very low intensity.

**SUGGESTED PHASE DIAGRAMS**

The suggested phase diagrams have been presented in Figs. 11 and 12. In the following we will explain how these diagrams have been constructed.

**Xe-N$_2$**

For $x=0.84$, the solid-fluid coexistence region is entered at 4.9 GPa, which is 0.3 GPa above the melting line of pure N$_2$.

In the coexistence region $v_F$ corresponds to the extrapolated frequency of the homogenous fluid. Therefore, $v_F$ is attributed to the fluid phase and, consequently, $v_{\delta}^\beta$ to the Raman mode of N$_2$ in the solid phase. There is no difference between the value of $v_{\beta}^\delta$ and the Raman frequency of pure $\beta$-N$_2$ (Fig. 3). This suggests that pure $\beta$-N$_2$ and the $\beta^*$ phase in the mixture are closely related. In the coexistence

![Graph](image1)

**FIG. 7.** (a) Typical $p$-$T$ scans for $x=0.35$ (circles). Lines are transition lines given in Fig. 8. (b) Simultaneous typical $v$-$T$ scans for $x=0.35$. Circles, Raman frequencies of solid phases; squares, Raman frequencies of fluid phase. Open symbols, $v$-$T$ scan at slightly higher pressure; solid symbols, $v$-$T$ scan at slightly lower pressure. Dashed and solid lines, Raman frequencies of pure $\delta$-N$_2$ at the corresponding temperature and pressure (Ref. 36); arrows, $S_{Ne^*}$ points for these two scans, measured by visual observations. Solid lines, low-pressure scan; dashed lines, high-pressure scan.

![Graph](image2)

**FIG. 8.** Comparison between spectra for $x=0.90$ and $x=0.35$.

![Graph](image3)

**FIG. 9.** An overview of the temperature scans for Ne-N$_2$. Circles, $S_3$-$S_2$-$\delta^*$ three-phase line; crosses, $S_2$-$\delta^*$-$F$ three-phase line; squares, $S_2$-$S_{Ne^*}$-$F$ three-phase line; triangles, $S_{Ne^*}F$ line for $x=0.35$. Solid lines, transition lines of pure N$_2$ (Refs. 17 and 24); dashed line, melting line of Ne (Ref. 37). Dotted and dash-dotted lines are fits through the points on the transition lines.

![Graph](image4)

**FIG. 10.** $v$-$T$ scan for $x=0.90$ at 11 GPa, compared with a scan for $x=0.35$ at the same pressure. Circles, $x=0.90$; crosses, $x=0.35$. Solid lines are the frequencies of pure $\varepsilon$- and $\delta$-N$_2$ (Ref. 36).
region, at the entrance point, the intensity of $\nu_F$ is already quite low, which can be explained by metastability. Since when the laser is focused on different spots in the sample space the intensity of $\nu_{S1}$ is always high compared to that of $\nu_F$, a considerable amount of $\beta^s$ must be present in the sample space. Therefore, a large amount of Xe dissolves into the $\beta^s$ phase, probably more than the amount of nitrogen that dissolves into solid Xe (see $x=0.075$).

At 6.0 GPa a phase transition occurs to the $\delta^s$ phase, since at 6.0 GPa, simultaneously $\nu_{S1}$ and the high-pressure modes $\nu_{1a,\delta^s}$, $\nu_{2a,\delta^s}$, and $\nu_{2b,\delta^s}$ were measured. There is no deviation between $\nu_{1a,\delta^s}$ and $\nu_{2a,\delta^s}$, and $\nu_1$ and $\nu_2$ of pure $\delta^s$-$N_2$. Therefore, we propose that this new phase ($\delta^s$) is a mixed solid of Xe dissolved into the $\delta^s$ phase of $N_2$. In the mixture, the $\beta^s$-$\delta^s$ transition is slightly shifted to lower pressures, since for pure $N_2$ the transition occurs at 6.2 GPa at 408 K. Because no other high-pressure phases have been detected for $x=0.84$, it is concluded that more than 16 mol% Xe dissolves into $\delta^s$-$N_2$. In contrast to $\delta^s$, the $\delta$ phase of pure Xe has only two Raman-active modes $\nu_1$ ($a$ sites) and $\nu_2$ ($c$ sites). Since the unit cell contains 3 times as much c sites as a sites, the intensity ratio of $\nu_1$ and $\nu_2$ is 0.33. For $x=0.84$ the ratio of the sum of the integrated intensities $\nu_{1a,\delta^s}$ and $\nu_{1b,\delta^s}$ and the sum of the integrated intensities $\nu_{2a,\delta^s}$ and $\nu_{2b,\delta^s}$ is 0.28±0.03. This suggests a preference of the Xe atoms for the $a$ sites in the mixed solid $\delta^s$.

For the mixture with $x=0.075$, the solid-fluid region is entered at 0.83 GPa, which is close to the estimated melting pressure of Xe from an extrapolation of the data of Ref. 25. In the solid-fluid coexistence region, the intensity of $\nu_F$ decreases as a function of pressure compared to that of $\nu_{Xe^s}$. Therefore, $\nu_F$ is attributed to the fluid phase and, consequently, $\nu_{Xe^s}$ to the Raman mode of $N_2$ in the mixed solid phase. It is concluded that $N_2$ dissolves into solid Xe (fcc). The rapid increase of $\nu_F$ as a function of pressure can be explained by a rapid change in composition of the fluid phase. This is reflected in the liquidus of Fig. 11. The absence of a fluid peak above 1.6 GPa suggests that (nearly) all the nitrogen dissolves into solid xenon in the range of 1.6 GPa $\leq p \leq 5.3$ GPa.

Above 5.3 GPa at least two phases must be present, since the intensity ratio of $\nu_{Xe^s}$ and $\nu_{S1}$ changes when the laser is focused at a different spot in the sample space. No discontinuity of $\nu_{Xe^s}$ has been found at 5.3 GPa; therefore, this peak still corresponds to $N_2$ in solid Xe. As a consequence, the mole fraction of $N_2$ in $S_1$ is larger than 0.075. The frequency $\nu_{S1}$ is considerably lower than the Raman frequencies $\nu_{S1}$ and $\nu_F$. Therefore, it is concluded that $S_1$ is different from these two phases. A stoichiometric compound is proposed for $S_1$ for the following reasons. Taking into account the intensity of the Raman peaks and the position of the other phase lines as proposed in Fig. 11, the mole fraction of xenon in this phase is probably between 0.4 and 0.7. The only candidate structures for a disordered solid of the pure components are the $\varepsilon$ phase of nitrogen (in pure nitrogen above 23 GPa) and the high-pressure intermediate close-packed phase in xenon (above 14 GPa at room temperature). We have shown already that less than 7.5% nitrogen dissolves into fcc xenon; thus, it is unlikely that in the closely related close-packed structure the solubility will be much higher. The $\varepsilon$ phase of $N_2$ is also not a good candidate for the following reasons. First, argon, which has nearly the same size as nitrogen, hardly dissolves into $\varepsilon$-$N_2$. Second, the transition line of the $\varepsilon$ phase must shift from 23 to 5.3 GPa. Third, the frequency of $\varepsilon^s$ must shift 4.6 cm$^{-1}$, while the frequency of $\delta^s$ is hardly shifted. Fourth, the Raman spectrum of $\varepsilon$-$N_2$ has two peaks, while only one peak is observed for $S_1$.

Ne-$N_2$

The first conclusion is that the solubility of nitrogen in solid neon is very low, if there is any solubility at all, as shown before. Then, since the spectrum of $\beta^s$ is only shifted by 0.2 cm$^{-1}$ compared to that of pure $\beta^s$-$N_2$, it is proposed that the $\beta^s$ phase is a disordered solid of neon dissolved into $\beta^s$-$N_2$. The $\beta^s$-$\delta^s$ transition is shifted by 1.4 GPa to higher pressures for $x=0.90$, and therefore neon dissolves into this phase. Since the $\beta^s$ phase coexists with a fluid phase for $x=0.90$, it is concluded that less than 10% neon dissolves into the $\beta^s$ phase.

When the pressure is increased, the amount of fluid decreases and, on the neon-rich side of the phase diagram, the concentration of the fluid phase shifts to higher nitrogen contents and vice versa as shown in Fig. 12. For $x=0.052$ and $x=0.90$, this should result in a rapid increase of the FWHM vs the pressure above the solid-fluid transition point, since the FWHM of the homogeneous fluid phase with $x=0.35$ is much higher than that with $x=0.052$ and $x=0.90$. This is indeed observed (Fig. 6). Moreover, the slope of $\nu_F$ versus the pressure should decrease for $x=0.052$ and $x=0.35$, and increase for $x=0.9$. This is observed as well (Fig. 5).

For the mixture with $x=0.052$, the mode $\nu_F$ disappears at the same pressure at which the modes of $S_2$ appear. Thus the mole fraction of $S_2$ is greater than 0.052. The mixture with $x=0.35$ shows both the Raman modes of $S_2$ and $\delta^s$, while for $x=0.052$ the modes of $\delta^s$ have not been measured. This proves that the mole fraction of $S_2$ is between 0.052 and 0.35. Because the frequencies of $S_2$ deviate from the Raman
modes of pure nitrogen and because the amount of neon that dissolves into the $\varepsilon$ and $\delta$ phases is less than 10% as will be explained below, $S_2$ must be a new solid phase. The peak positions and the integrated intensity ratio of $\nu_{52,a}$ and $\nu_{52,b}$ have the same value for the two mixtures mentioned above; thus, it is proposed that $S_2$ is a van der Waals compound.

The fact that the modes of $\delta^*$ have the same frequency as pure $\delta$-$N_2$ (Fig. 5) strongly suggests that the structure of $\delta^*$ is very similar to $\delta$-$N_2$. However, the decrease of the intensity ratio of $\nu_{1,\delta^*}$ and $\nu_{2,\delta^*}$ from 0.33 for pure $N_2$ to 0.24 $\pm$ 0.03 for the mixtures suggests that the neon atoms dissolve into $\delta$-$N_2$ and that they are mainly positioned at the $a$ sites. Considering the intensity ratio, the neon mole fraction of $\delta^*$ must be 0.05 or higher. Since also for $x=0.90$ the Raman frequencies of $\delta^*$ and $S_2$ have been measured simultaneously in the scan of $\nu$ versus $T$, it is concluded that there is a coexistence region of $\delta^*$ and $S_2$ and that, therefore, the nitrogen mole fraction of $\delta^*$ is more than 0.90 in this region.

Because the intensity of the modes of $\delta^*$ changes with respect to that of $S_3$ when the laser is focused at a different spot in the sample space, it is concluded that $\delta^*$ and $S_3$ are two different solid phases. Since the spectrum of $S_3$ is different from that of the various solid phases of pure $N_2$, it can be either a van der Waals compound or a disordered solid of neon dissolved into $N_2$. As shown below, it cannot be a disordered solution of neon in $\varepsilon$-$N_2$. Although another disordered solid cannot be excluded, the van der Waals compound seems the most likely candidate because of the following arguments. The spectrum of $S_3$ is very similar to that of the van der Waals compound He(N2)$_{11}$, as measured by Olijnyk and Jephcoat$^{16}$ and Scheerboom$^{17}$. Analogous to the spectrum of the compound He(N2)$_{11}$, the spectrum of $S_3$ shows three modes up to 15 GPa, $\nu_{53,a}$ with about the same frequency as $\nu_1$ of $\varepsilon$-$N_2$, $\nu_{53,b}$ with a frequency slightly higher than $\nu_2$ of $\varepsilon$-$N_2$, and $\nu_{53,c}$ with a frequency slightly lower than $\nu_2$ as shown in Fig. 5. In addition, also similar to the spectra of the compound He(N2)$_{11}$, the intensity ratio of $\nu_{53,a}$ and $\nu_{53,b}$ is reduced compared to that of $\nu_1$ and $\nu_2$ of pure $\varepsilon$-$N_2$, and the mode $\nu_{53,b}$ shows an asymmetry on the high-frequency side (marked by arrows in Fig. 8). There are no candidate structures for a disordered solid in neon, while in nitrogen there are only indications for a new phase $\eta$ above 20 GPa at high pressure,$^{27}$ but this could not be confirmed in x-ray studies.$^{28}$ Most probably, $\eta$ is only a small modification of the $\varepsilon$ phase.

For $x=0.35$, a coexistence region of $S_2$ and $S_3$ has been found. Since the mole fraction of $S_2$ is less than 0.35, the mole fraction of $S_3$ must be more than 0.35. For $x=0.90$ no coexistence region of $S_1$ with $S_3$ has been observed, indicating that the mole fraction of $S_3$ is less than 0.9. For $x=0.9$, a coexistence region is proposed of $S_1$ with $\varepsilon^*$ and $\delta^*$, respectively. In the scan with $x=0.9$ of $\nu$ versus the temperature, up to 196 K the mode $\nu_{1,\varepsilon^*}$ has been found indeed, with the same frequency as $\nu_1$ of pure $\varepsilon$-$N_2$. The Raman mode $\nu_2$ of $\varepsilon$-$N_2$ has the same frequency as $\nu_{53,b}$ within experimental accuracy, which explains why this mode could not be detected. Since the two modes $\nu_{53,b}$ and $\nu_{2,\varepsilon^*}$ cannot be distinguished, the intensity ratio between $\nu_{1,\varepsilon^*}$ and $\nu_{2,\varepsilon^*}$ cannot be determined, and therefore it is impossible to draw conclusions on the solubility of Ne in $\varepsilon$-$N_2$. At 212 K the mode $\nu_{1,\varepsilon^*}$, observed at 196 K and lower temperatures, is no longer observed in the temperature scan for $x=0.9$. Therefore, it is proposed that for these temperatures $S_1$ co-exists with $\delta^*$. In pure $N_2$, the $\varepsilon$-$\delta$ transition occurs at 208 K at this particular pressure, and thus, within experimental accuracy, the $\varepsilon^*$-$\delta^*$ transition is not shifted for the mixture. Because $S_3$ has been found in coexistence with $\varepsilon^*$ as well as with $\delta^*$, it cannot be a disordered solid of neon dissolved in the $\varepsilon$ or $\delta$ phase of nitrogen. Since the frequencies $\nu_{53,a}$ and $\nu_{53,b}$ are the same as $\nu_1$ and $\nu_2$ of $\delta$-$N_2$, respectively, it is impossible to distinguish between these four peaks. The intensity ratio of $\nu_{53,a}$ and $\nu_{53,b}$ and that of $\nu_{53,c}$ and $\nu_{53,b}$ changes when the laser is focused at a different spot in the sample space, and in addition the ratio is higher than for $x=0.35$. These observations are in agreement with a $\delta^*$-$S_3$ coexistence region, since the intensity ratio for $\delta^*$ is higher than that of $S_3$. This also explains the behavior of the spectra as a function of pressure for $x=0.90$ at and above 11.0 GPa at room temperature. Depending on the position of the laser spot, in the sample space we have found the modes of $S_3$, the modes of $\delta^*$, or both.

**DISCUSSION AND CONCLUSIONS**

The high-pressure phase diagrams of Xe-$N_2$ and Ne-$N_2$ show a rich variety of solid phases. The present measurements show that the ratio of the diameters of the molecules, $\alpha$, is only a very rough indication of the phase behavior of binary molecular systems. There is no solubility of the large nitrogen molecules in solid neon ($\alpha=0.74$), in agreement with theoretical predictions on hard-sphere systems. On the other hand, the small neon molecules do dissolve into the $\beta^*$ and $\delta^*$ phases of $N_2$, even though the diameter ratio $\alpha$ is much smaller than 0.85. This results in a considerable deviation of the Hume-Rothery rule. The system Xe-$N_2$ shows indeed mutual solubility, as expected on the basis of the diameter ratio ($\alpha=0.89$), but in contrast to the computer simulations on hard-sphere systems, there is a larger solubility of the larger xenon molecules in nitrogen than vice versa. Therefore, the van der Waals interaction, the anisotropy, ori-
entational degrees of freedom, and the Coulomb forces probably play an important role in the mutual solubility in molecular solids at high pressure.

As discussed before, the Ne and Xe atoms dissolve substitutionally into $\delta$-N$_2$. Analogous to the system Ar-N$_2$, the Ne and Xe atoms have a preference for the $a$ sites. The intensity of the $v_1$ vibron is more reduced in the system Ne-N$_2$ than in the system Xe-N$_2$, although the maximum solubility is in the order of 10% for Ne-N$_2$ and more than 16% for Xe-N$_2$. Thus the Ne atoms have a stronger preference for the $a$ sites than the Xe atoms.

In this work it was found that for Ne-N$_2$ the $\delta^+\beta^*$ transition is shifted to higher pressures, just like Ar-N$_2$, while for Xe-N$_2$ the $\delta^+\beta^*$ transition is shifted to slightly lower pressures. For the system Ne-N$_2$ the $\epsilon-\delta$ transition is not shifted, in contrast to He-N$_2$ and Ar-N$_2$ where it is shifted to lower and higher pressures, respectively.

In Xe-N$_2$ one van der Waals compound $S_1$ has been found, while in Ne-N$_2$ ($\alpha=0.74$) at least two compounds exist: $S_2$ and $S_3$. The spectrum of $S_2$ consists of two distinct peaks, suggesting that the N$_2$ molecules are positioned on at least two distinct sites. The spectrum of $S_1$ is very similar to that of He(N$_2$)$_{11}$. Analogous to the interpretation of Olijnyk and Jephcoat $^\dagger$ for He(N$_2$)$_{11}$, a close relation is proposed between the structure of $S_1$ and $\epsilon$-N$_2$ and the Ne atoms may substitute or displace the N$_2$ molecules on the $a$ sites.

Previously, the following compounds have been found in molecular systems: $A_1B$ (N$_2$-He,$^\dagger$ $\alpha=0.62$), $A_2B$ (CH$_4$-H$_2$, $\alpha=0.72$), $AB$ [CH$_4$-H$_2$ (Ref. 9)] $AB_2$ [He-He,$^\dagger$ $\alpha=0.83$; Ar-H$_2$, $\alpha=0.80$; CH$_4$-H$_2$ (Ref. 9)] and $AB_4$ [CH$_4$-H$_2$ (Ref. 9)], with $B$ the smaller molecule. In computer simulations and analytical theories on hard-sphere systems, only $AB$, $AB_2$, and $AB_3$ were found to be stable for 0.2 $\leq\alpha\leq0.42$, 0.42 $\leq\alpha\leq0.59$, and 0.54 $\leq\alpha\leq0.61$, respectively. $AB_3$ has not been found in molecular systems, and although $AB_2$ and $AB$ have been detected, the structure which has been determined experimentally for all compounds is different from the stable structure for hard-sphere systems. Because of the large discrepancy between hard-sphere and molecular systems, it is concluded that compound formation in molecular systems is not only determined by hard-sphere packing considerations. Just as in the case of disordered solids, the van der Waals interaction, the anisotropy, orientational degrees of freedom, and the Coulomb forces probably play an important role in the compound formation at high pressure.

Free energy calculations $^{29,31,32}$ with realistic potentials in the local harmonic approximation show that the compound Ne(He)$_2$ with the Laves structure is stable. Contrary to the hard-sphere systems, in the free energy calculations, the energy is the driving force for the compound formation. $^{32}$ Therefore, it can be concluded that good potentials are needed to understand the formation of van der Waals compounds.

The mole fraction of $S_1$ in the system Xe-N$_2$ is most probably 0.4 $\leq x \leq 0.7$. Therefore, the most plausible structures for $S_1$ in Xe-N$_2$ are (N$_2$)Xe with $x=0.5$ and (N$_2$)$_2$Xe, with $x=0.67$. The mole fraction of $S_2$ in the system Ne-N$_2$ is 0.052 $\leq x \leq 0.35$. Therefore, we suggest that $S_2$ has the structure N$_2$(Ne)$_2$ with $x=0.33$, N$_2$(Ne)$_4$ with $x=0.2$, or N$_2$(Ne)$_{13}$ with $x=0.07$. The mole fraction of $S_3$ is probably 0.6 $\leq x \leq 0.9$, and the similarity of the spectrum with that of He(N$_2$)$_{11}$ suggests that the number of N$_2$ molecules in the unit cell is large compared to the number of Ne atoms. The real structures can only be obtained with x-ray diffraction.

The frequency shift as a function of concentration can be explained for both systems by the following arguments. In the fluid phase the frequency increases at increasing Ne and decreasing Xe concentrations, respectively, due to the change in dispersive contribution. $^{33}$ In the solid phase, in addition to the dispersive contribution, the structure influences the shift. In contrast to the fluid phase, in the $\delta^+$ phase neither for Xe-N$_2$ nor for Ne-N$_2$ has a shift been detected for the frequency with respect to pure $\delta$-N$_2$, while the solubility of Xe in this phase is larger than 16 mol %. This may partly be explained by the peculiar structure of $\delta$-N$_2$. The Xe and Ne atoms have a preference for the $a$ sites. Since the first neighbors of molecules on both the $a$ sites and $c$ sites are $c$ sites, the nearest neighbors of N$_2$ molecules in the mixed solids are N$_2$ molecules. Preliminary simulations $^{34}$ show that in this case the dispersive contribution is not very sensitive to the composition. The $\epsilon$ phase is obtained from the $\delta$ phase by a small distortion. Therefore, the same arguments might explain why the frequency of $S_3$ in Ne-N$_2$ is only slightly shifted as compared to that in pure $\epsilon$-N$_2$. In Xe-N$_2$, the mode $\nu_{\delta^+\epsilon^+}$, with a mole fraction lower than 0.075, has a much lower frequency than that in pure N$_2$. The decrease in the frequency is so large that it cannot be attributed only to a change in the dispersive contribution. In Ref. 35 it is shown that because Xe crystallizes on a larger grid than N$_2$, the axial force exerted on the N$_2$ molecules in solid Xe is lower than that in pure N$_2$, resulting in an additional decrease in the frequency.

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34 J. P. J. Michels (private communication).

