Temperature fitting of partially resolved rotational spectra

Citation for published version (APA):
https://doi.org/10.1088/1748-0221/7/02/C02054

DOI:
10.1088/1748-0221/7/02/C02054

Document status and date:
Published: 01/01/2013

Document Version:
Publisher’s PDF, also known as Version of Record (includes final page, issue and volume numbers)

Please check the document version of this publication:
- A submitted manuscript is the version of the article upon submission and before peer-review. There can be important differences between the submitted version and the official published version of record. People interested in the research are advised to contact the author for the final version of the publication, or visit the DOI to the publisher's website.
- The final author version and the galley proof are versions of the publication after peer review.
- The final published version features the final layout of the paper including the volume, issue and page numbers.

Link to publication

General rights
Copyright and moral rights for the publications made accessible in the public portal are retained by the authors and/or other copyright owners and it is a condition of accessing publications that users recognise and abide by the legal requirements associated with these rights.

- Users may download and print one copy of any publication from the public portal for the purpose of private study or research.
- You may not further distribute the material or use it for any profit-making activity or commercial gain
- You may freely distribute the URL identifying the publication in the public portal.

If the publication is distributed under the terms of Article 25fa of the Dutch Copyright Act, indicated by the “Taverne” license above, please follow below link for the End User Agreement:
www.tue.nl/taverne

Take down policy
If you believe that this document breaches copyright please contact us at:
openaccess@tue.nl
providing details and we will investigate your claim.

Download date: 18. Apr. 2021
Temperature fitting of partially resolved rotational spectra

This article has been downloaded from IOPscience. Please scroll down to see the full text article.

2012 JINST 7 C02054
(http://iopscience.iop.org/1748-0221/7/02/C02054)

View the table of contents for this issue, or go to the journal homepage for more

Download details:
IP Address: 131.155.110.104
The article was downloaded on 11/04/2012 at 16:45

Please note that terms and conditions apply.
Temperature fitting of partially resolved rotational spectra

AFH van Gessel,\textsuperscript{a,1} B. Hrycak,\textsuperscript{b} M. Jasiński,\textsuperscript{b} J. Mizeraczyk,\textsuperscript{b} J.J.A.M. van der Mullen\textsuperscript{a} and P.J. Bruggeman\textsuperscript{a}

\textsuperscript{a}Eindhoven University of Technology, Department of Applied Physics, P.O. Box 513, 5600 MB, Eindhoven, The Netherlands
\textsuperscript{b}Centre for Plasma and Laser Engineering, The Szewalski Institute of Fluid-Flow Machinery, Polish Academy of Sciences, Fiszera 14, 80-231 Gdańsk, Poland

E-mail: a.f.h.v.gessel@tue.nl

ABSTRACT: In this paper we present a method to automatically fit the temperature of a rotational spectrum. It is shown that this fitting method yields similar results as the traditional Boltzmann plot, but is applicable in situations where lines of the spectrum overlap. The method is demonstrated on rotational spectra of nitric oxide from an atmospheric pressure microwave plasma jet operated with a flow of helium and air, obtained with two different methods: laser induced fluorescence and optical emission spectroscopy. Axial profiles of the rotational temperatures are presented for the ground NO X state and the excited NO A state.

KEYWORDS: Data processing methods; Plasma diagnostics - interferometry, spectroscopy and imaging

\textsuperscript{1}Corresponding author
1 Introduction

The rotational spectrum of a species is a valuable data source for determining gas temperatures inside a plasma. The most common method is a Boltzmann plot, where the peak intensities of the rotational lines are plotted logarithmically against the energy of the rotational transition. Since the line intensities follow a Boltzmann distribution, the slope of the points is equal to $1/k_B T$, where $k_B$ is Boltzmann’s constant, and $T$ the temperature [1]. Several methods exist based on Boltzmann plots to determine temperatures with the intensities of one, two of more rotational lines [2–4].

A Boltzmann plot is insightful, because deviations from the Boltzmann distribution are easily observed if points deviate from the linear fit [5]. Furthermore the method does not require a lot of computational power. The difficulty arises in the determination of the peak intensities. Since this is usually done by hand, or by taking the local maxima of a spectrum. This requires that the width of the instrumental profile of the measurement device is much smaller than the spectral distance between the rotational lines. Furthermore, when the number of data points per peak is small, a fitting procedure needs to be applied to obtain the peak maxima. This can also induce additional uncertainties.

Overlapping spectra can lead to inaccuracies in Boltzmann plots. Especially for molecules with relative high mass such as NO and N$_2$ the rotational lines are close together — typically in the order of a few pm — as the energy difference between the levels increases with decreasing reduced mass of the rotator. This puts high demands on the used spectrometer. Partially overlapping spectra can even occur in LIF measurements as the FWHM of dye lasers is typically of the order of 2 pm.

In cases where Boltzmann plots are not applicable, temperature determination is often done using software packages that calculate rotational spectra. Examples of such software are Lifbase from Luque et al. [6], Specair from Laux [7], and Lifsim from Bessler et al. [8]. These programs perform calculations of spectra similar as shown in this paper. However, the temperature is always an input parameter, and these programs do not yield automatically the best fit to a measured
spectrum. In literature several examples exist of temperatures that are determined by fitting. For example for NO a method is derived from Lifsim by Bessler et al. [9]), for N$_2^+$ a fitting method is presented by Linss et al. [10], and for N$_2$, O$_2$, OH and NO by Andre et al. [11]. However finding a temperature is still often done manually by fitting the spectrum by eye. This is a subjective method, and leads to serious inconsistencies often found in scientific papers on the accuracy of the gas temperature, which is often overestimated.

In the next section we propose a method to determine the rotational temperature automatically by fitting a spectrum directly to the measured spectrum with a least-square method. The instrumental profile is incorporated into the fitting method, making it possible to fit spectra with overlapping lines, lines from different rotational branches, and implement a background correction. In section 4 we apply our method to rotational spectra of nitric oxide (NO), obtained with Optical Emission Spectroscopy (OES) and Laser Induced Fluorescence (LIF).

2 Fitting method

Temperature determination is most often made by rotational spectra obtained from OES and LIF. We describe briefly the procedure below:

2.1 OES

For OES the light emitted by the plasma is measured with a spectrometer. For temperature determination we only need the relative intensities, and we consider rotational states from a single vibrational band. This means that we only have to consider factors which depend on the rotational state. The intensity $I_i$ of a line of transition $i$ is reduced to,

$$I_i \propto A_i \cdot (2J_i + 1)e^{-\frac{E_i}{kT}} \quad \text{(OES)} \quad (2.1)$$

where $A_i$ is the Einstein emission coefficient for transition $i$. $J_i$ and $E_i$ represent the rotational number and energy of the emitting (upper) state of the transition $i$. The $2J_i + 1$ factor is to account for the degeneracy of the rotational states, where $J_i$ is the rotational quantum number of emitting (upper) state of the transition $i$. The temperature dependence is included in a Boltzmann exponent, which means that this method is physically equivalent to a Boltzmann plot. $E_i$ is the energy of the rotational state, given by,

$$E_i = B_v \cdot J_i(J_i + 1) \quad (2.2)$$

where $B_v$ is the rotational constant. Higher order terms represent a fraction less then 1% for $J_i < 40$ for most species [12], and can thus be neglected in many cases.

2.2 LIF

In the case of LIF the spectrum is measured by scanning a dye laser along rotational transitions. The fluorescence signal is often a broadband signal as a spectrometer with open slit or an interference filter (typically passing 10 nm) is used, such that all rotational transitions from one vibrational band are detected.
LIF is a two-step process: first a photon is absorbed, followed by emission of the populated level. After Uddi et al. [13] the intensity $I_i$ of a line of transition $i$ can be written as,

$$I_i \propto \frac{A_{v''}^{v'}}{Q + \sum_{v'} A_{v''}^{v'} B_i} \cdot (2J_i + 1) e^{-\frac{E_i}{k_B T}}$$  \hspace{1cm} (2.3)$$

where $B_i$ is the Einstein absorption coefficient for transition $i$, induced by the laser. $A_{v''}^{v'}$ is the Einstein emission coefficient for the detected vibrational transition, divided by the sum over all vibrational states, plus the quenching coefficient $Q$. Since this $A$ coefficient is independent of the transition $i$, equation (2.3) can be reduced to,

$$I_i \propto B_i \cdot (2J_i + 1) e^{-\frac{E_i}{k_B T}} \hspace{1cm} (LIF)$$  \hspace{1cm} (2.4)$$

This equation is valid if: 1. The laser intensity is low and there is no saturation, i.e. stimulated emission can be neglected; 2. Vibrational energy transfer is negligible, or at least independent of the rotational state; 3. Rotational energy transfer is faster than the lifetime of the excited state, such that there is a redistribution of the rotational levels of the excited state; 4. Quenching is independent of the rotational state; 5. The plasma is optically thin, i.e. the absorbed laser energy is negligible compared to the total laser energy.

To calculate the total spectrum $S(\lambda)$, the line intensities have to be multiplied with a line profile $g(\lambda)$ and summarized. We also add a constant background $C$,

$$S(\lambda) = C + \sum_i I_i \cdot g(\lambda - \lambda_i)$$  \hspace{1cm} (2.5)$$

The line profile is an arbitrary function, composed of the spectral distribution function of the transition, and the laser profile (in case of LIF) or the instrumental profile of the spectrometer (in case of OES). In most cases the line profile can be a adequately approximated by a Voigt profile,

$$g(\lambda) = G(\lambda, \Delta_G) \otimes L(\lambda, \Delta_L)$$  \hspace{1cm} (2.6)$$

where $G$ is a Gaussian curve with width $\Delta_G$ and $L$ is a Lorentzian curve with width $\Delta_L$.

Of the above quantities $A_i$, $B_i$, $\lambda_i$, $J_i$ and $B_i$ are species properties, which are published for many species. Assuming these quantities are known, the resulting spectrum function is $S(\lambda; I_0, T, C, \Delta_G, \Delta_L)$. This function can be programmed into a computer and fitted to a measured spectrum, with $\lambda$ as independent parameter and $I_0$, $T$, $C$, $\Delta_G$ and $\Delta_L$ as fitting parameters. We used MATLAB to perform the fitting, in particular the function FIT, which is part of the CURVE FITTING TOOLBOX. This function performs a least-square fit, and has the ability to calculate confidence intervals on the fitting parameters. It is possible to determine $C$, $\Delta_G$ and $\Delta_L$ with other experimental methods, which reduces the number of fitting parameters to two.

3 Experimental setup

The fitting method described in the previous section is applied rotational spectra of nitric oxide (NO), obtained with OES and LIF.
The NO source is an atmospheric pressure microwave plasma jet, as described by Hrycak et al. [14] (see figure 1). The input microwave power is 30 W. The plasma is operated with a flow of a 6.0 slm helium mixed with 0.2 slm of air, resulting in a flow speed of approximately 1 m/s. The tube ends in ambient air.

The OES measurements are performed using a 1 m Jobin Yvon spectrometer with a SBIG CCD camera, with an instrumental profile width of approximately 25 pm, and a spectral distance between pixels of 6 pm. The light is collected using a lens and a fiber, which provides a spot size in the plasma of 2 mm diameter. The measured spectrum is around 247 nm, which corresponds to the transition NO $A^2\Sigma^+ - X^2\Pi$ ($v = 0 - 2$).

The LIF measurements are performed with a Sirah dye laser, pumped with an Edgewave Nd:YAG laser at 355 nm and a repetition rate of 4 kHz. The dye laser beam is frequency doubled, resulting in a UV beam around 226 nm, and 10 µJ per pulse, with a line width of 1.4 pm. The laser excites the NO $X^2\Pi - A^2\Sigma^+$ ($v = 0 - 0$) transition. The detection system consists of a McPherson EUV monochromator and a Hamamatsu R8486 photomultiplier connected to a counting system. The monochromator is set at 247 nm with a wide exit slit of 1 mm and a FWHM of 10 nm, such that all rotational transitions from the NO $A^2\Sigma^+ - X^2\Pi$ ($v = 0 - 2$) vibrational transition are detected. The laser beam is not focussed, in order to avoid saturation. The size of the laser beam and the optics that focus the LIF signal onto the entrance slit of the monochromator make that the measurements have a detection volume of approximately 1 mm$^3$.

4 Applying the fitting method

Figure 2 shows an example of a LIF spectrum of NO. To determine the rotational temperature we applied the fitting method described in section 2. The rotational energy transfer rate for the NO $A$ state is in the order $10^{10}$ cm$^3$s$^{-1}$ for helium [15] and air [16]. For atmospheric pressure RET is in the order of 1 ns, while the measured decay time of the excited state is 48 ns, thus the assumption
Figure 2. Rotational spectrum of NO, obtained with LIF inside the plasma jet at 3 mm above the tube end. The theoretical spectrum is plotted, with a fitted values $T_{\text{rot}} = 860 \pm 43$ K, $\Delta_G = 2.2 \pm 0.3$ pm and $\Delta_L = 1.0 \pm 0.2$ pm.

Figure 3. Boltzmann plot of a LIF measurement of NO inside the plasma jet at 3 mm above the tube end. The used lines are indicated with $\times$ in figure 2. The linear fit gives a temperature of $T_{\text{rot}} = 846 \pm 106$ K.

that the rotational states are redistributed in the excited state is justified. Vibrational transfer is much slower [17], and can be neglected.

From Lifbase [6, 18] we obtained the parameters $B_i$ (for the NO $X - A (v = 0 - 0)$ transitions), $\lambda_i$ and $J_i$ (for the NO $X (v = 0)$ state) for each line in the spectrum. Furthermore $B_v = 1.6961 \text{ cm}^{-1}$ [12]. The fitted spectrum is shown in figure 2, with a fitted temperature of $T_{\text{rot}} = 860 \pm 43$ K. The error margin is a 95% confidence interval as provided by the MATLAB \textsc{fit} function.

From the same LIF measurement lines from the P$_2$ branch and the overlapping Q$_{12}$ branch are used to make a Boltzmann plot (see figure 3). The corresponding temperature is calculated to be $T_{\text{rot}} = 846 \pm 106$ K (95% confidence interval).

Results from the OES measurement are shown in figure 4. In this case the resolution of the spectrum is insufficient to completely resolve the rotational lines. Consequently, making
Figure 4. Rotational spectrum of NO, obtained with OES from inside the plasma jet at 3 mm above the tube end. The theoretical spectrum is plotted, with a fitted values $T_{rot} = 1829 \pm 11$ K, $\Delta G = 24.6 \pm 0.8$ pm and $\Delta L = 8.1 \pm 0.5$ pm.

Figure 5. Axial temperature profiles obtained with different methods.

A Boltzmann plot is not possible. The temperature fitting, however, is still applicable. The parameters $A_i$ (for the NO $A - X$ ($v=0$-2) transitions), $\lambda_i$ and $J_i$ (NO $A$ ($v = 0$)) are obtained from Lifbase [6, 18], and $B_v = 1.9862$ cm$^{-1}$ [12]. The resulting temperature is $T_{rot} = 1829 \pm 11$ K (95% confidence interval).

Similar OES and LIF measurements are performed at different axial positions in and above the plasma. The axial temperature profiles are shown in figure 5. The temperatures from the Boltzmann plot and the fitting method correspond within the margin of error, where the fitting method has the smallest error margins. While the LIF signal from the NO $X$ ground state is still measurable in the afterglow of the plasma, the emission from the NO $A$ excited state can only be measured inside the plasma.

The OES temperatures from the NO $A$ excited state are higher than the LIF temperatures from the NO $X$ ground state. Larger rotational temperatures than the gas temperature for NO $A$ is common, see for example Staack et al. [19]. This can be explained by the formation process (most probably the association process) causing specific large rotational excitation, while due to
the limited rotational life time of the excited state equilibration with the gas temperature is not possible. In the case of ground state NO, the lifetime is much larger and thus the NO $X$ rotational temperature is a better estimate of the gas temperature.

5 Conclusions

We present a method to automatically fit a rotational temperature to a measured spectrum. Because the line positions and Einstein coefficients are taken from literature, the method is relatively simple. The temperature fitting works well in situations where a Boltzmann plot is not applicable due to overlapping lines. We applied the method to rotational spectra of NO in an atmospheric pressure plasma jet. The fitted temperatures are $T_{rot} = 860 \pm 43$ K for NO in the $X^2\Pi$ ground state (measured with LIF), and $T_{rot} = 1829 \pm 11$ K for NO in the $A^2\Sigma^+$ excited state.

To obtain MATLAB scripts that implement the described fitting method, please contact the authors.

References


