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Citation for published version (APA):

Document status and date:
Published: 01/01/2015

Document Version:
Accepted manuscript including changes made at the peer-review stage

Please check the document version of this publication:
• A submitted manuscript is the version of the article upon submission and before peer-review. There can be important differences between the submitted version and the official published version of record. People interested in the research are advised to contact the author for the final version of the publication, or visit the DOI to the publisher’s website.
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• The final published version features the final layout of the paper including the volume, issue and page numbers.

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C/T/E Energy Storage - 09

Shrinking Core Model for the Reaction-Diffusion Problem in Thermo-Chemical Heat Storage

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Keywords: Thermochemical heat storage, Shrinking core model, Reaction diffusion interaction, Lithium sulphate monohydrate, Thermal dehydration

SUMMARY

In this work, we develop a kinetic model to study the dehydration reaction of Li₂SO₄.H₂O single particles involving interaction between the intrinsic chemical reaction and the bulk diffusion. The mathematical framework of the model is based on the shrinking core model. A variable-grid, finite-difference method with fully implicit formulation is used for solving the model. It is found that the Damköhler number \( Da = (k_r t_0) / (D_e c_0) \) plays an important role in determining the nature of the diffusion/reaction dynamics. A very small \( Da \) value means that the overall reaction is controlled by the intrinsic chemical reaction at the interface, while a very large \( Da \) value means that the overall reaction is controlled by the diffusion of water through the product phase. Moreover, the numerical results of fractional conversion calculated in the model are in good agreement with the theoretical analysis under extreme cases in which either diffusion (large \( Da \)) or reaction (small \( Da \)) dominates the dehydration process. With consideration of numerical solutions at various \( Da \) values, it is concluded that both intrinsic reaction and mass diffusion are important in determining the reaction kinetics within a range of \( Da \) values between 0.1 and 10.

INTRODUCTION

Dehydration reactions of salt hydrates are of prime importance in various industrial and environmental applications such as thermochemical heat storage. An adequate understanding of the mechanism and kinetics of the reactions is essential to analyse and then to optimize the reaction rate. In this work, the thermal dehydration of lithium sulphate monohydrate (see Eq. 1) is investigated.

\[
\text{Li}_2\text{SO}_4\cdot\text{H}_2\text{O}(s) \leftrightarrow \text{Li}_2\text{SO}_4(s) + \text{H}_2\text{O}(g) \tag{1}
\]

This reaction used to be recommended by the Kinetics Committee of the International Confederation of Thermal Analysis (ICTA) as a material for the study of the reaction kinetics of solid-state reaction. Therefore, its reaction kinetics has been extensively studied using various kinetic models, but significant differences in the Arrhenius parameters have been reported (Brown et al. 1992; Brown et al. 1993; Ferchaud et al. 2012; Kanungo and Mishra
One of the main reasons is the use of different methods of data analysis. It is well
known that the complicated process of dehydration is composed of several stages, e.g., the
intrinsic reaction of bond breaking, formation of nuclei and their growth, diffusion of gaseous
products, and heat transfer. The fact is that, in general, the reaction kinetics is not to be
determined by a single mechanism but by a combination of several mechanisms.
Unfortunately, kinetic models for the dehydration rate are often simplified to a certain
mechanism due to the ease of application. Kinetic studies based on them do not necessarily
provide reliable kinetic data.

![Illustration of the shrinking core model](image)

We want to consider the kinetics of the dehydration reaction in a more fundamental
mechanistic level, where the overall reaction is dependent on intrinsic reaction and mass
diffusion. The idea of the shrinking core (sharp interface) model is used to describe the
thermal dehydration of \( \text{Li}_2\text{SO}_4\cdot\text{H}_2\text{O} \) salt hydrates. As shown schematically in Fig. 1, the
dehydration process of a salt hydrate crystal includes a phase change only at the sharp
interface of the core. The particle size does not change but the radius of the core decreases as
the reaction proceeds. In the fully hydrated state, the salt hydrate crystal is in the \( \alpha \) phase. The
dehydration process begins when water vapor at the surface of the particle starts to escape
from the crystal structure. The transformation of the salt hydrate \( (\alpha \) phase) at the surface
promotes the formation of the dehydrated phase \( (\beta \) phase). Next, the adsorbed water that is
lost at the surface of the particle is replenished by diffusion of water molecules from the salt
hydrate \( \alpha \) phase to the surface through the \( \beta \) phase. The dehydration process is complete when
all the water from the salt crystal has been lost and the particle consists only of the \( \beta \) phase.

The shrinking core model was first developed by Yagi and Kunii (Yagi and Kunii 1955; Yagi
and Kunii 1961) to describe the reactions between gaseous species and solid particles, and has
been developed to incorporate various physicochemical phenomena under a variety of
conditions (Knorr et al. 2012). A summary of previous work on this model can be found in
the article of Gupta and Saha (Gupta and Saha 2003). There are a few applications of this
model to a solid-state reaction where the interaction between the interface reaction and mass
diffusion is involved simultaneously. One example is about the thermal decomposition of
\( \text{SrCO}_3 \) compacts, but the calculation is limited to the pseudo-steady state (Arvanitidis et al.
1999). In other fields of application such as the electrochemistry of the discharging of a metal
hydride electrode (Deshpande et al. 2011; Subramanian et al. 2000), metallurgy of the phase
transformation (Schuh 2000) and hydrogen storage (Saetre 2006), only the diffusion of
gaseous products is considered without reaction at the moving interface. Recently, Cui et al.
(Cui et al. 2013) presented a particularly complete model to describe the interface-reaction
controlled diffusion in binary solids, such as \( \text{Li}_x\text{Si} \), which can account for the finite
def ormation kinetics, diffusion-stress interaction and the interplay between bulk diffusion and
intrinsic chemical reaction.
In this paper a shrinking core model involving the intrinsic chemical reaction and bulk diffusion is developed. A variable-grid, finite-difference method with fully implicit formulation is used for solving the governing equations of the model. Numerical solutions under various situations regarding the competition between intrinsic reaction and mass diffusion are presented for discussion.

**NUMERICAL MODEL**

The dehydration of a spherical salt hydrate particle is assumed to follow the sequence described in Fig. 1. The mass flux of the reversible reaction considered is written as:

\[ J_{H,O} = k_r (1 - \frac{c_r}{c_{eq}}) \]  

(2)

where \( k_r \) is the reaction rate constant at the interface, \( c_r \) is the water concentration at the interface and \( c_{eq} \) is the equilibrium concentration. In terms of the propagation of the reaction interface position \( dr_c \) in a time interval \( dt \), the mass flux of water is given by:

\[ J_{H,O} = -c_0 \frac{dr_c}{dt} \]  

(3)

where \( r_c \) is the interface position. Combining Eq. 2 and Eq. 3, we have:

\[ \frac{dr_c}{dt} = -k_r \frac{1}{c_0} \left( 1 - \frac{c_r}{c_{eq}} \right), \quad r = r_c \]  

(4)

As the particle is dehydrated, the core of the fully hydrated material (\( \alpha \) phase) shrinks. The concentration distribution \( c(r) \) of water in the \( \beta \) phase is governed by Fick’s second law:

\[ r^2 \frac{\partial c}{\partial t} = D_e \left( r^2 \frac{\partial}{\partial r} \left( r^2 \frac{\partial c}{\partial r} \right) \right), \quad r_c \leq r \leq r_0 \]  

(5)

where \( r \) is the spatial variable and \( D_e \) is the effective diffusivity of water through the porous \( \beta \) phase.

In order to preserve the mass balance at the interface, the mass flux relative to the moving interface must be considered. The amount of water generated from the interfacial reaction must be equal to the amount transported away from the interface. In particular, mass flux of convection must be converted from a fixed frame to the moving interface with an absolute velocity \( dr_c / dt \). So the boundary condition at the moving interface is given as:

\[ k_r \left( 1 - \frac{c_r}{c_{eq}} \right) = -D_e \frac{\partial c(r,t)}{\partial r} \bigg|_{r=r_c} - \frac{dr_c}{dt} c_r, \quad r = r_c \]  

(6)

Replacing \( dr_c / dt \) and rearranging the equation, the boundary condition at the interface can be rewritten as:
The boundary condition at the outer surface of the particle is given as:

\[ c(r, t) = c_g, \quad r = r_0 \]  

where \( c_g \) is the water concentration in the environment.

For the model to be as general as possible, the following nondimensional spatial and temporal variables are defined:

\[ \hat{r} = \frac{r}{r_0}; \quad \hat{t} = \frac{D r}{r_0}; \quad \hat{c} = \frac{c - c_g}{c_{eq} - c_g}; \quad \hat{c}_g = \frac{r}{r_0} \]

The corresponding dimensionless forms of the above equations are:

\[ \frac{d\hat{r}}{dt} = -Da(1 - \hat{c}), \quad r = r_c \]  

\[ r^2 \frac{\partial \hat{c}}{\partial \hat{t}} = \frac{\partial}{\partial \hat{r}} (r^2 \frac{\partial \hat{c}}{\partial \hat{r}}), \quad r_c \leq r \leq 1 \]  

with boundary conditions

\[ \frac{\partial \hat{c}(r, t)}{\partial \hat{r}} \bigg|_{\hat{r}=0} = -Da(1 - \hat{c})(M - \hat{c}), \quad r = r_c \]  

\[ c(r, t) = 0, \quad r = 1 \]  

where the hats are dropped for convenience, \( \lambda = \frac{(c_{eq} - c_g)}{c_{eq}}, \) \( M = \frac{(c_0 - c_g)}{(c_{eq} - c_g)} \) and \( Da = \frac{(k_r r_0)}{(D_r c_0)} \) is the Damköhler number expressing the ratio of the chemical reaction rate to the mass transport rate.

The present model is solved by numerical scheme proposed by Illingworth and Golosnoy (Illingworth and Golosnoy 2005). The solution is based on the finite difference method with fully implicit formulation. The moving interface is tracked by a variable grid method using a Landau transformation so that the moving interface is fixed at a grid point. A considerable advantage of this scheme is that finite difference equations are derived in such a way that mass conservation during the calculation is ensured.

**RESULTS AND DISCUSSIONS**

In this section, the shrinking core model is solved for a spherical particle for various values of the Damköhler number. It is worth noting that in all calculations, only the reaction rate constant \( k_r \) is tuned in order to achieve a range of \( Da \) values. The water concentration profile
as a function of dimensionless time for various \( Da \) values is shown in Fig. 2. Each profile shows the water content in the \( \beta \) phase from the moving interface to the particle surface. The arrow in Fig. 2 indicates the direction of propagation of the sharp interface between \( \alpha/\beta \) phases. The shape of each profile at a certain time level is attributed to the interaction between the intrinsic reaction and the mass diffusion. From the definition of the Damköhler number, a small \( Da \) value means that the interfacial chemical reaction rate is relatively low compared to the diffusion rate of water molecules in the \( \beta \) layer. In the case of \( Da = 0.1 \), the interface moves very slow compared to the diffusion of water in the crystal. As a result, water released by reaction at the interface has sufficient time to drain away. In contrast, for a large \( Da \) value, e.g. \( Da = 100 \), the intrinsic chemical reaction is faster, and water cannot escape from the particle quickly enough. This leads to an increase of the water concentration near the interface, which in turn diminishes the propagation of the reaction interface. In other words, the Damköhler number plays a significant role in determining the nature of the diffusion/reaction dynamics. A very small \( Da \) value means that the overall reaction is controlled by the intrinsic chemical reaction at the interface, while a very large \( Da \) value means that the overall reaction is controlled by the diffusion of water through the \( \beta \) phase.

![Figure 2: Profiles of water concentration distribution in a spherical particle for various \( Da \) values.](image)

Another remarkable feature in the present results is that the water concentration at the interface decreases after it reaches a peak. The explanation is that as a structural model, the rate of reaction is essentially related to the reaction area which changes during the course of transforming the spherical particle. As the reactive area shrinks, the mass production at the interface also declines. As a result, a large percentage of gaseous water can diffuse away from
the particle during the same time period, which in return will accelerate the chemical reaction at the interface.

To further investigate the limiting mechanism of the reaction kinetics between the intrinsic reaction and the bulk diffusion process, the dimensionless interface position and normalized dimensionless interface velocity as a function of dimensionless time for the different values of $Da$ are plotted in Fig. 3. At low values of $Da$, the interface position in Fig. 3(a) exhibits a linear dependence on time, which is consistent with the zero-order kinetics model used in the intrinsic reaction. In contrast, at high values of $Da$ the process is diffusion controlled, where the plot of the interface position against time is curved.

![Figure 3: (a) Dimensionless interface position for various Da values; (b) Dimensionless interface velocity for various Da values (right).](image)

In Fig. 3(b), the normalized dimensionless interface velocity is shown, which is defined by:

$$v_c = -\frac{1}{Da\lambda} \frac{dr}{dt} = 1 - c_r.$$  

As discussed above, the velocity is constant at a small $Da$ value corresponding to a linear $r_c$ profile. With the increase of the $Da$ value, the dimensionless velocity decreases rapidly, which is attributed to the diffusion limitation. In the case of $Da = 1$, the transition from a reaction to a diffusion controlled process can be observed. The overall kinetics of the dehydration reaction is initially controlled by the interfacial chemical reaction during which the interface velocity is constant. As the interface moves away from the particle surface, water vapor cannot diffuse out of the crystal efficiently, which leads to the accumulation of water molecules. Consequently, the movement of interface position is slowed down gradually because the propagation of the reaction front is proportional to the difference of water concentration from the equilibrium concentration (see Eq. 4). The shape factor discussed above also explains the increase of the interface velocity in Fig. 3(b) after it reaches a minimum value. It is interesting to note that the interface velocity for $Da = 1$ almost increases to its initial value at the end of the reaction. The reaction kinetics will again be governed by the intrinsic reaction after it is governed by the intrinsic reaction in the initial stage of the dehydration process and by the mass diffusion during the second stage.

By integrating the water concentration, the fractional conversion of the salt hydrate can be plotted as a function of dimensionless time for different values of $Da$, see Fig. 4. Besides the numerical results at various $Da$ values in solid lines, the line marked by ‘+’ is the solution from the zero-order reaction model and the marked line by ‘o’ is the analytical solution of a pure diffusion process. As expected, our calculation results compare well with the theoretical analysis under the extreme cases in which either diffusion (large $Da$) or reaction (small $Da$)
dominates the dehydration process. In the case of $Da = 1$, the shape of fractional conversion is nearly linear as in the case of $Da = 0.1$. However, it is a result of the combined influence of interfacial reaction, mass diffusion and particle shape, instead of pure interfacial reaction as in $Da = 0.1$. Based on the previous discussion, it is safe to say that the range of $Da$ values for which both intrinsic reaction and mass diffusion are important in determining the reaction kinetics is between 0.1 and 10, outside of which only a single mechanism, either the intrinsic reaction or the mass diffusion, is dominant in the reaction kinetics.

![Figure 4: Fractional conversion as a function of dimensionless time for various values of Da during dehydration of a spherical particle. The marked lines by ‘o’ and ‘+’ are analytical solutions for pure diffusion and zero-order reaction, respectively.](image)

**CONCLUSIONS**

In this work a shrinking model is developed for describing the dehydration reaction of $Li_2SO_4.H_2O$ single particles. It includes the interaction between the intrinsic chemical reaction and the mass diffusion simultaneously. The Damköhler number is found to distinguish the controlling step during the transformation. A very small $Da$ value means that the overall reaction is controlled by the intrinsic chemical reaction at the interface, while a very large $Da$ value means that the overall reaction is controlled by the diffusion of water in the product phase. Fractional conversion calculated from the model agrees with the theoretical analysis under the extreme cases in which either diffusion (large $Da$) or reaction (small $Da$) dominates the dehydration process. Within the range of $Da = 0.1$ and $Da = 10$, both the intrinsic reaction and the mass diffusion are important in determining the reaction kinetics. In future work, experiments are designed to validate the model and to examine the dehydration reaction in a much more fundamental way. Information gained on the mechanisms of the dehydration reaction of $Li_2SO_4.H_2O$ monocrystals could form the basis for the development of a more rigorous kinetic model.

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