Scalable anti-Markovnikov hydrobromination of aliphatic and aromatic olefins†

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To improve access to a key synthetic intermediate we targeted a direct hydrobromination-Negishi route. Unsurprisingly, the anti-Markovnikov addition of HBr to estragole in the presence of AIBN proved successful. However, even in the absence of an added initiator, anti-Markovnikov addition was observed. Re-examination of early reports revealed that selective Markovnikov addition, often simply termed “normal” addition, is not always observed with HBr unless air is excluded, leading to the rediscovery of a reproducible and scalable initiator-free protocol.

Terminal alkenes are readily converted into valuable synthetic intermediates for metal-mediated cross-coupling reactions by hydro-metallation to give organo-boron† and other organometallics. These reactions typically proceed under steric control to give the primary organometallic, often designated as the “anti-Markovnikov” product, a term that refers back to seminal work done over 140 years ago by Victor Markovnikov on the analogous addition of HI to alkenes.

We recently applied such a hydrometallation-Suzuki approach to the synthesis of bromopyridine 2, a key intermediate in the synthesis of mechanically chiral rotaxanes, stabilised reactive organometallic species and interlocked catalysts; hydroboration of commodity chemical estragole (1a) with 9-BBN-H followed by an in situ cross coupling with 2,6-dibromopyridine yielded 2 in a concise manner (Scheme 1). However, on scale up we encountered problems with purification due to the borinic acid by-product, which, in addition to the high cost of 9-BBN-H, led us to explore other routes to 2. Accordingly, we explored a Negishi approach employing an organozinc species produced in situ from bromide 3a, itself accessed in three steps from cheap and readily available hydroxyphenyl propionic acid. However, although the Negishi coupling step is efficient and scalable, the three-step synthesis of bromide 3a once again proved cumbersome on scale up.

These issues led us to consider the direct anti-Markovnikov hydrobromination of estragole to produce 3a in order to combine the key advantages of both syntheses. This approach proved extremely successful giving rapid access to 3a and thus 2 in multi-gram quantities. More importantly, as a result of these studies we made an initially surprising observation: even in the absence of added initiators the hydrobromination of 1a proceeds in reasonable selectivity to give the anti-Markovnikov product.

Here we report how this observation led to the rediscovery of simple scalable conditions for synthesis of primary bromides under “initiator free” conditions from alkyl and aryl alkenes. Our results increase the availability of primary bromides directly from feedstock alkene substrates.

The hydrobromination of olefins is generally held to proceed through two competing pathways: polar pathway I via the most stable carbocation typically resulting in the branched, Markovnikov product, and radical pathway II via the most stable radical, resulting in the linear, anti-Markovnikov product (Fig. 1). To favour pathway II, reactions are carried out in apolar solvents in the presence of radical initiators (the “peroxide effect”) or under irradiation. The Markovnikov and anti-Markovnikov products are also often simply called the “normal” and “abnormal” products respectively.

Surprisingly, direct synthesis of primary bromides from monosubstituted alkenes by reaction with HBr appears to be a

relatively under-used reaction,12 a simple search gave only 330 examples compared with the >48 000 such bromides reported.13 We were also surprised to be unable to find anything recognisable as an organic methodological study in which a variety of substrates are screened under the same conditions, presumably because most work on the peroxide effect was carried out in the first half of the 20th century with each paper reporting only a few examples.11 Thus, most recent reports of this transformation are confined to isolated examples as part of a larger synthetic campaign.

A brief screen of conditions14,15 identified the use of HBr in PhMe in the presence of AIBN as appropriate, giving 3a in excellent 97 : 3 selectivity (Scheme 2).16 A minor drawback of this procedure on larger scales is the relatively high loading (13 mol%) of AIBN required. Unfortunately, attempts to reduce this led to erratic results (see below). However, the excess AIBN could be removed readily simply by filtering the reaction mixture through silica prior to evaporation and applying this procedure allowed us to reliably produce 3a across a range of scales (1–200 mmol) in excellent yield (98%).

Moreover, these conditions proved general for representative monosubstituted aliphatic alkenes (1b–f). The slightly reduced selectivity in the case of allyl ether 1e may be due to anomiceric assistance by the proximal O atom favouring the linear product. To our knowledge there are no previous reports of this transformation are confined to isolated examples as part of a larger synthetic campaign.

Based on these results the reaction of HBr in toluene with AIBN appears general for aliphatic alkenes but only applicable to electron-stabilised alkenes bearing strongly electron withdrawing substituents. However, during our attempts to reduce the AIBN loading we made an unexpected observation: on small scales (1 mmol), even when no external initiator was added a significant selectivity for primary bromide 3a was still observed, albeit with poor reproducibility. Based on the received wisdom of undergraduate chemistry this result is superficially surprising as, in the absence of added initiators or irradiation, the Markovnikov product is predicted in systems that lack significant electronic bias.19

In order to understand this observation we returned to the early publications in the field, in particular an excellent contemporary review from Walling.11 This revealed a number of interesting points often omitted in recent discussions. Firstly, many of the early investigations of the addition of HBr to alkenes were conducted using the neat alkene, rather than under dilute conditions where the polar pathway is dramatically retarded. Secondly, in order to observe the Markovnikov product, great care was always taken to use extremely pure alkene substrates and exclude oxygen and other adventitious oxidants because, although Markovnikov addition is the rule for HCl and HI, the case of HBr is far more nuanced; even in the presence of vanishingly small quantities of oxidants, anti-Markovnikov addition can compete and even dominate in the case of alkenes that are not activated to Markovnikov addition.

Thus our surprise observation was in fact common knowledge when the peroxide effect was first discovered. Perhaps unhelpfully, although Markovnikov’s rule is often discussed in the context of hydrobromination, HBr is the only hydrohalic acid in which this outcome is sometimes hard to observe experimentally as anti-Markovnikov addition often competes due to the presence of adventitious oxidants. Indeed, the terms “normal” and “abnormal” addition actually seem to have originally referred to the reactions of HI and HCl in which no peroxide effect is observed and thus the abnormal addition actually refers to the contrast with these products, rather than with that observed in the case of HBr “normally”.

During our literature search, one of the early examples of initiator-free hydrobromination caught our attention. Sherrill and co-workers reported in 1934 that HBr added as a solution in AcOH gave reliable anti-Markovnikov addition to pent-1-ene and hept-1-ene in hexane.20 Sherrill’s conditions have been applied only twice in synthesis and the origin of the unusual reaction outcome was not commented upon.21,19a,b These conditions are particularly attractive as the use of a commercially available solution of HBr in AcOH is operationally simpler than using HBr gas to produce an HBr solution.

Under Sherrill’s original conditions, aliphatic alkenes 1a–f, with the exception of allyl ether 1e, were hydrobrominated in excellent selectivity, comparable to that observed in toluene in the presence of AIBN. Furthermore, these conditions deliver improved selectivity (>80%) for the primary bromide product in the case of styrene itself (1g) and p-fluoro-styrene (1h). Furthermore, even weakly (1i, 1l) electron withdrawing para sub-

**Scheme 2** Addition of HBr in PhMe to alkenes. Figures in parentheses refer to the selectivity 3–5. Reagents and conditions: HBr in PhMe (sat.), AIBN (13 mol%), 0 °C, 2 h.
stition was obtained predominantly through a spontaneous hydrobromination reaction of monosubstituted terminal aliphatic and, for the first time, aromatic alkenes. The omission of initiators such as AIBN or benzoylperoxide removes the need for purification of the products, allowing them to be taken forward directly in further synthetic manipulations. To be clear, we achieved this not by discovering new conditions but by investigating and generalising a previously reported but largely forgotten procedure from Sherrill and co-workers. That this procedure has remained largely ignored for so long is surprising given its synthetic utility and may be in part due to the counterintuitive nature of the conditions, in that no obvious initiator is added, combined with the lack of previous methodological investigations. The results presented here should increase the synthetic availability of primary bromides as synthetic intermediates derived from feedstock monosubstituted terminal alkenes.

Acknowledgements

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Notes and references


3 (a) W. Markownikoff, *Ann. der Chemie und Pharm.*, 1870, 153, 228; (b) V. V. Markovnienoff, *C. R. Hebd. Seances Acad. Sci.*, 1875, 81, 668.

4 It has been suggested that Markownikov’s “rule” was the result of inspired guess-work rather than significant experimental evidence. For an interesting account see: P. Hughes, *J. Chem. Educ.*, 2006, 83, 1152.


11 For an excellent contemporary review of the early work on the addition of hydrohalous acids to alkenes and the peroxyde effect, including the observation of anti-Markownikov addition in the absence of added initiators, see: F. R. Mayo and C. Walling, *Chem. Rev.*, 1940, 27, 351.


13 A search using the Reaxys database with the criteria below provides 330 reactions when limited to those that use HBr as a reactant. By comparison, a search for primary bromides of this form suggest >46 000 such compounds are known.

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\begin{align*}
\text{G} & \rightarrow \text{G}^+ + \text{Br}^-
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15 Unfortunately, although the desired conversion of estragole to 3a has previously been reported, detailed experimental conditions were not provided: J. Delobelle, M. Fetixon, P. Baranger, J. Schalbar and M. J. Trefouel, *C. R. Hebd. Seances Acad. Sci.*, 1957, 244, 2402.

16 The majority of literature procedures call for HBr gas, readily produced by reaction of Br₂ with tetraele, to be passed through the reaction mixture with all other components present. However, we found this led to poor reproducibility, presumably due to poor control over the concentration of HBr. Pre-saturation of the solvent with HBr prior to addition of the substrate removed this limitation. Furthermore, the solution of HBr in toluene could be stored for up to one month in the freezer.

17 Kharasch and co-workers reported the hydrobromination of styrene in dilute pentane solution with debenzyol peroxide to give an 80:20 ratio in favour of the primary bromide. Unfortunately, detailed conditions were not provided: (a) C. Walling, M. S. Kharasch and F. R. Mayo, *J. Am. Chem. Soc.*, 1939, 61, 2693. The only other reliable report of the direct hydrobromination of a styrene derivative to give the primary bromide involved pentachloro or tetrachloro-benzenes which have significant electronic bias towards to the primary product.¹³


