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Cooperative Folding of Linear Poly(dimethyl siloxane)s via Supramolecular Interactions

Marcin L. Śleżczkowski, E. W. Meijer,* and Anja R. A. Palmans*

The synthesis and characterization of graft copolymers are reported based on linear poly(dimethyl siloxane) (PDMS) and chiral, pendant benzene-1,3,5-tricarboxamides (BTAs). The copolymers differ in degree of polymerization (DP) and BTA graft density. Characterization of the bulk materials at room temperature reveals that the BTAs aggregate in a helical fashion via threefold hydrogen-bond formation within the PDMS matrix. A significant degree of hydrogen bonding persists up to 180 °C, regardless of DP and BTA content. Analysis of the solution behavior by 1H NMR spectroscopy indicates that BTA aggregation occurs in CDCl3, a solvent normally suppressing aggregation. Circular dichroism (CD) spectroscopy in 1,2-dichloroethane shows strong CD effects and reveals that increasing the DP and decreasing the BTA graft density results in an increase in the cooperativity of the BTA aggregation. Dynamic light scattering indicates the formation of particles with sizes of 400 nm. This is the first time that polymers with pendant BTAs show a sharp transition between a nonaggregated and aggregated state, a behavior similar to the one observed for “free” BTAs. The cooperative aggregation is attributed to the strong phase-segregation between the PDMS backbone and the BTAs, in combination with a high propensity of these polymers to form multichain aggregates.

1. Introduction

Poly(dimethyl siloxane)s (PDMS) are a highly interesting and versatile class of polymers.1–3 Their simple chemistry combined with the propensity to phase segregate4 as well as the unique flexibility of the siloxane backbone5–7 provides outstanding possibilities in the bottom-up synthesis of functional materials that found applications in microfluidic devices,8 actuators,9 as dielectric insulators in electronic devices,10 Their biocompatibility with human tissue11 and living cells12 can potentially extend the applications toward biomaterials. Previous studies showed that strong phase-segregation is a powerful force that can drive the self-assembly of (macro)molecules in the bulk phase. Covalent block copolymers of poly(dimethyl siloxane) and poly(ethylene oxide),13,14 poly(lactic acid),15–17 poly(methacrylate),18 poly(styrene),19–21 and poly(2-vinylpyridine)22 phase segregate to form well-ordered morphologies in the bulk. In addition, oligo- and poly(dimethyl siloxanes) have been combined with supramolecular hydrogen-bonding motifs to afford supramolecular block copolymers,23 thermoplastic elastomers,24,25 self-healing elastomers26–27 and liquid crystalline materials.28 The introduction of functionality in supramolecular phase segregating systems has been achieved by direct symmetrical end-functionalization of oligo(dimethyl siloxane) (oDMS) of discrete lengths with ureidopyrimidinone (UPy) units that are able to dimerize via fourfold hydrogen bond formation.29 Protected UPy-oDMS conjugates exhibited liquid crystalline properties whereas deprotection caused a sharp transition toward block copolymer-like behavior. In recent years, we and others have evaluated in detail the folding of single polymer chains driven by pendant hydrogen-bonding units in water30–35 and organic media.36–43 Notably, benzene-1,3,5-tricarboxamides (BTAs) have been found to efficiently fold polymers chains into compartmentalized structures in water. In organic media, in contrast, there is a propensity for multichain aggregation when dynamic hydrogen-bonding motifs are applied.33 Often, these studies used poly(methacrylate)- or poly(norbornene)-based backbones in poor solvents, in which the conformational flexibility of the polymer backbone is reduced. We wondered in how far the high flexibility of the PDMS backbone, in combination with helical self-assembling BTAs, could enhance effective chain folding in organic media. In addition, we previously established that the folding of polymers with BTA pendants is noncooperative,23,36 which is in stark contrast to the highly cooperative nature of the self-assembly of free BTAs.44 The origin of the noncooperative behavior was attributed to the formation of several domains in which BTAs were aggregated within one polymer particle. This lack of ability of all BTAs to aggregate into one helical stack was associated with the high entropic penalty of the polymeric backbone to fold around a BTA stack.10 The entropic penalty of the folding of the polymeric backbone offsets the normal cooperative behavior of BTAs, which has an enthalpic origin. The use of a highly flexible backbone, in contrast, could result in cooperative folding of the polymer chains.
Inspired by the work of Bouteiller and co-workers,[45] we here decorate a PDMS backbone with BTA units affording graft polymers, encoded as PDMS-g-BTA. We vary the polymer length and BTA density along the backbone. The materials obtained are studied in detail in bulk using variable temperature infrared spectroscopy (VT-IR), circular dichroism (CD) spectroscopy, small angle X-ray scattering (SAXS), polarized optical microscopy (POM), and differential scanning calorimetry (DSC), whereas in dilute solution CD and dynamic light scattering (DLS) are used. Here, we show that strong phase segregation between the BTAs and the PDMS backbone in solution results in unparalleled cooperativity in the folding of the polymers, but at the same time also in the formation of large particles.

2. Results and Discussion

The synthesis of PDMS-g-BTA polymers (Figure 1A) required two building blocks, namely poly(dimethyl siloxane-co-methylhydrosiloxane) (PDMS-co-PHMS) copolymers and a BTA-olefin. PDMS-co-PHMS is commercially available with different average degrees of polymerization and molar PHMS content, which determines the degree of functionalization. The enantiomerically pure (S)-BTA-olefin has been synthesized in three steps via a parallel synthesis with good yield (29%) and high purity starting from trimesic acid according to the synthetic route shown in Schemes S1 and S2 in the Supporting Information.

PDMS-g-BTA was synthesized via direct hydrolylation of three commercially available PDMS-co-PHMS with average molecular weights of 25 kg mol\(^{-1}\) (5% PHMS), 6 kg mol\(^{-1}\) (8% PHMS), and 2 kg mol\(^{-1}\) (16% PHMS). The BTA olefin was coupled to the three polymers applying a platinum(0)-1,3-di-vinyl-1,1,3,3-tetramethyldisiloxane complex (the Karstedt catalyst). All polymers were purified by means of BioBeads in tetrahydrofuran (THF), which afforded a good separation between final product and the residual BTA-olefin impurities. The three polymers were fully characterized by \(^1\)H NMR and size exclusion chromatography (SEC) in THF (Figures S1–S3, Supporting Information). The results are summarized in Table 1.

As an illustrative example, the \(^1\)H NMR (THF-d\(_4\)) and SEC trace (THF) of P1-g-BTA are shown in Figure 1B,C. The \(^1\)H NMR spectrum of P1-g-BTA in THF-d\(_4\) reveals the characteristic signals of the BTA core at 8.38 (Ph-H) and 7.98 ppm (NH) as well as the aliphatic protons (between 3.40 and 0.80 ppm). The largest peak at 0.10 ppm corresponds to the PDMS methyl protons. The BTA functionalization of PDMS is unambiguously confirmed by presence of a triplet at 0.5 ppm which corresponds to the methylene group attached to the PDMS backbone. The degree of BTA functionalization was quantified by comparing the integrals of the signals at 0.1 and 0.5 ppm (see the Supporting Information for details). For P1-g-BTA this results in a BTA graft density of 3%. P2-g-BTA and P3-g-BTA show higher graft densities of 9% and 14%, respectively (Table 1).

The SEC trace of P1-g-BTA in THF (Figure 1C) shows a unimodal peak, corresponding to an \(M_n\) of 42.1 kg mol\(^{-1}\) and a molar mass dispersity \(D\) of 2.05. The measured \(M_n\) is higher than expected (38 kg mol\(^{-1}\) based on the SEC data of PDMS-co-PHMS and calculated degree of functionalization), presumably a result of the graft polymer topology in combination with the fact that the SEC column is calibrated with polystyrene standards. Analogously, P2-g-BTA and P3-g-BTA show also higher than expected values for \(M_n\).

Interestingly, the differences in \(M_n\) and BTA graft density are directly reflected in the physical appearance of the copolymers: whereas P1-g-BTA (\(M_n = 42.1\) kg mol\(^{-1}\), 3 mol% BTAs) is a flexible solid, P2-g-BTA (\(M_n = 17.6\) kg mol\(^{-1}\), 9 mol% BTAs) and P3-g-BTA (\(M_n = 10.7\) kg mol\(^{-1}\), 14 mol% BTAs) are relatively hard and brittle materials. To investigate how the number of BTA grafts in combination with the degree of polymerization (DP) affects the bulk material properties, we performed Fourier-transform infrared spectroscopy (FT-IR), CD spectroscopy, DSC, SAXS, and POM studies. In the FT-IR spectrum of P1-g-BTA (Figure 2A) absorptions bands characteristic for helically aggregated BTAs[46] at 3238, 1641, and 1562 cm\(^{-1}\) are observed as well as bands corresponding to the siloxane backbone (1250, 1010, and 790 cm\(^{-1}\)).[47] Bands at similar positions were found for P2-g-BTA and P3-g-BTA (Figures S4 and S5, Supporting Information), indicating that irrespective of the physical appearance of the polymers, helically aggregated BTAs are present in the

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### Table 1. Characterization data of polymers PDMS-g-BTA.

<table>
<thead>
<tr>
<th>Polymer</th>
<th>(M_n)(^{a}) (kg mol(^{-1}))</th>
<th>(D)(^{b})</th>
<th>mol% BTA(^{a})</th>
<th>(n)(^{c})</th>
<th>(m)(^{d})</th>
<th>(g)(^{e})</th>
</tr>
</thead>
<tbody>
<tr>
<td>P1-g-BTA</td>
<td>42.1</td>
<td>2.05</td>
<td>3</td>
<td>439</td>
<td>14</td>
<td>−0.0012</td>
</tr>
<tr>
<td>P2-g-BTA</td>
<td>17.6</td>
<td>1.66</td>
<td>9</td>
<td>123</td>
<td>12</td>
<td>−0.0019</td>
</tr>
<tr>
<td>P3-g-BTA</td>
<td>10.7</td>
<td>3.08</td>
<td>14</td>
<td>57</td>
<td>9</td>
<td>−0.0023</td>
</tr>
</tbody>
</table>

\(^{a}\)Determined by SEC in THF, calibrated with polystyrene standards; \(^{b}\)Determined by \(^1\)H NMR; \(^{c}\)\(n\) is the number of siloxane units within the polymer backbone, which was calculated using the methylhydroxiloxane content provided by the supplier and \(M_n\); \(^{d}\)\(m\) is the number of BTA units within the polymer backbone, which was calculated using \(M_n\) and mol% BTA; \(^{e}\)The anisotropy value \(g\) was determined from a film spin-coated on a quartz substrate.

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**Figure 1.** A) Chemical structure of PDMS-g-BTA. B) \(^1\)H NMR spectrum of P1-g-BTA in THF-d\(_4\). \(^{a}\) Peaks at 1.73 and 3.57 correspond to THF; peak at 2.46 ppm corresponds to water. C) SEC trace of P1-g-BTA in THF.
bulk samples. This helical aggregation was further confirmed with CD spectroscopy. Since the chiral BTA grafts all possess (S)-stereogenic centres, we expect that an M-helical sense is favored in the formed aggregates, which is characterized by a negative CD effect.[48] Films of all polymers were spin coated on quartz and all PDMS-g-BTA polymers exhibited a negative CD effect with an extremum at 225 nm (Figure 2B). The anisotropy value \( \Delta \varepsilon \) was calculated (Table 1) and the absolute values increase in the series P1-g-BTA to P3-g-BTA. This suggests that although helical aggregation is present in all graft copolymers, the degree to which this occurs depends on the density of the BTA grafts.

To further unravel the mesoscopic organization of PDMS-g-BTA, POM and SAXS measurements were performed. Under crossed polarizers, birefringence was observed in all samples (Figure S6, Supporting Information). However, the undefined textures indicate a poorly ordered structure. This lack of long range order was confirmed by the SAXS profile in which only small, broad peaks were visible in the low \( q \) region (Figure S7, Supporting Information).

The stability of the BTA aggregates of PDMS-g-BTA polymers in bulk was studied using POM, DSC, and VT-IR. POM showed a loss of birefringence. From the above, we conclude that the lack of long range order was confirmed by the SAXS profile in which only small, broad peaks were visible in the low \( q \) region (Figure S7, Supporting Information).

Temperature-dependent CD measurements were conducted by cooling the dilute solutions of PDMS-g-BTA from 100 to 0 °C (Figure 3B). At 90 °C, no CD effect is visible, indicating that the BTA are not aggregated. Upon cooling, an abrupt increase in the CD effect is observed for P1-g-BTA at 72 °C, a less abrupt increase takes place in case of P2-g-BTA at 70 °C, and P3-g-BTA
exhibits a cooling curve similar to previously reported for carbon-based polymers with BTA pendants, which all exhibited a gradual increase of the CD effect.\(^{[30,31,35,37,38,49]}\) The shapes of the cooling curves are connected to the increase of the density of BTA pendants. Despite the differences in the sharpness of the transition between P2-g-BTA and P3-g-BTA, the values of $|\Delta \varepsilon|$ are similar. P1-g-BTA shows a cooling curve similar to the ones recorded for free BTAs in alkane solvents, which is indicative for highly cooperative BTA aggregation.\(^{[44]}\) As the degree of functionalization is increased, the sharpness in the transition, and therefore the cooperativity, diminishes. In addition, increasing the BTA concentration of P1-g-BTA from 50 to 200 $\mu$M shows that the $T_c$ remains at around 72 °C, the CD effect increases proportionally with the BTA concentration, and that the cooperativity of the transition remains very high (Figure 3C).

DLS experiments performed on P1-g-BTA in DCE at 20 °C show the presence of particles with sizes of around 400 nm hydrodynamic diameter with a low dispersity (Figure 3D), indicating that the formation of the particles is thermodynamically controlled.

Based on the results above, we propose that the remarkable cooperative BTA aggregation observed in P1-g-BTA is related to the highly incompatible nature of PDMS and the covalently attached BTAs combined with the formation of helical BTA aggregates. In addition to the phase-segregation process, there is also the intrinsic flexibility of the PDMS backbone\(^{[4,7]}\) which significantly reduces the entropic penalty of backbone folding in comparison to the more stiff PMMA backbones, permitting a high degree of aggregation between pendant BTAs. Also, the poor solubility of PDMS in DCE at lower temperatures could play an important role. At room temperature, the mixture of the pure precursor polymer PDMS-co-PHMS in DCE is turbid at 10 mg mL\(^{-1}\) whereas the mixture becomes clear when heated above 50 °C. Covalently attaching the BTAs changes the solvent compatibility of the graft copolymer.

At the same time, the PDMS backbone may shield the BTAs from the solvent, hereby enhancing hydrogen-bonding interactions. Although it appears from the above results that the collapse of the PDMS backbone in DCE creates a confined space that promotes BTA aggregation, the exact nature of the interaction triangle between DCE, BTA, and PDMS is not fully clear and is topic of further investigations.

3. Conclusions

We have shown that supramolecular assembly of BTAs attached to PDMS leads to phase segregated structures in bulk. The formation of helical aggregates stabilized by threefold hydrogen bonding was confirmed by IR and CD spectroscopy. The formed superstructures were stable up to 180 °C. In addition, we demonstrated that proper selection of the polymeric backbone can lead to extraordinary folding properties. P1-g-BTA exhibited highly cooperative assembly of the BTAs in DCE, which was not achieved for any BTA-graft copolymers before and suggests that this unusual behavior is connected to the synergy between BTA-backbone, BTA-solvent, and backbone-solvent interactions. This combination leads to the thermodynamically controlled formation of multichain aggregates in which nearly all BTAs are aggregated. The degree of functionalization in combination with the degree of polymerization also has a significant impact on the folding characteristics, which indicates that a proper balance in BTA density is needed. An important question is how sequence and dispersity control will affect the cooperative folding behavior of PDMS-g-BTA and further studies will be devoted to this topic.

Supporting Information

Supporting Information is available from the Wiley Online Library or from the author.

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Conflict of Interest
The authors declare no conflict of interest.

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