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Interplay between Surface Chemistry, Precursor Reactivity, and Temperature Determines Outcome of ZnS Shelling Reactions on CuInS$_2$ Nanocrystals

Anne C. Berends,† Ward van der Stam,‡† Jan P. Hofmann,‡ Eva Bladt,§ Johannes D. Meeldijk,⊥ Sara Bals,§ and Celso de Mello Donega*‡†

†Condensed Matter and Interfaces, Debye Institute for Nanomaterials Science, Utrecht University, Post Office Box 80000, 3508 TA Utrecht, The Netherlands
‡Laboratory of Inorganic Materials Chemistry, Department of Chemical Engineering and Chemistry, Eindhoven University of Technology, Postbox 513, 5600 MB Eindhoven, The Netherlands
§EMAT, Department of Physics, University of Antwerpen, Groenenborgerlaan 171, 2010 Antwerpen, Belgium
⊥Electron Microscopy Utrecht, Debye Institute for Nanomaterials Science, Utrecht University, 3584 CH Utrecht, Netherlands

Supporting Information

ABSTRACT: ZnS shelling of I–III–VI$_2$ nanocrystals (NCs) invariably leads to blue-shifts in both the absorption and photoluminescence spectra. These observations imply that the outcome of ZnS shelling reactions on I–III–VI$_2$ colloidal NCs results from a complex interplay between several processes taking place in solution, at the surface of, and within the seed NC. However, a fundamental understanding of the factors determining the balance between these different processes is still lacking. In this work, we address this need by investigating the impact of precursor reactivity, reaction temperature, and surface chemistry (due to the washing procedure) on the outcome of ZnS shelling reactions on CuInS$_2$ NCs using a seeded growth approach. We demonstrate that low reaction temperatures (150 °C) favor etching, cation exchange, and alloying regardless of the precursors used. Heteroepitaxial shell overgrowth becomes the dominant process only if reactive S- and Zn-precursors (S-ODE/OLAM and ZnI$_2$) and high reaction temperatures (210 °C) are used, although a certain degree of heterointerfacial alloying still occurs. Remarkably, the presence of residual acetate at the surface of CIS seed NCs washed with ethanol is shown to facilitate heteroepitaxial shell overgrowth, yielding for the first time CIS/ZnS core/shell NCs displaying red-shifted absorption spectra, in agreement with the spectral shifts observed for type-I band alignment. The insights provided by this work pave the way toward the design of improved synthesis strategies to CIS/ZnS core/shell and alloy NCs with tailored elemental distribution profiles, allowing precise tuning of the optoelectronic properties of the resulting materials.

The search for Pb- and Cd-free nanocrystals (NCs) has greatly intensified in recent years. Copper indium sulphide (CIS) in particular is one of the most extensively investigated alternative materials since CIS NCs display tunable photoluminescence (PL) through the visible and near-infrared (NIR) spectral window, large absorption cross-sections across a broad spectral range, and low toxicity. These properties make colloidal CIS NCs promising materials for a variety of applications (viz., light-emitting diodes, photovoltaics, bioluminescence, and spectral converters in displays and luminescent solar concentrators). However, the PL quantum yield (QY) of CIS NCs is typically low (<10%) due to exciton trapping at surface defects, which facilitates nonradiative decay. To realize the full potential of CIS NCs, it is thus imperative that strategies to boost their PL QYs are developed. The extensive research carried out over the last decades on the prototypical II–VI semiconductor NCs has clearly established that exciton trapping at surface defects can be effectively prevented by overgrowth of shells of wider band gap materials, thereby leading in recent years to CdSe-based core/shell quantum dots (QDs) with near unity ensemble PL QYs. The excellent quality of II–VI core/shell QDs motivated the scientific community to develop similar strategies for CIS NCs. To date, these efforts have mainly focused on the overgrowth of ZnS shells since ZnS is a nontoxic, stable, and abundant material, with a small lattice mismatch with respect to CIS (viz., 2%), and a wide band gap (3.54 eV) with (bulk) band...
edges positioned such that a type-I band alignment with CIS is expected.\textsuperscript{6,13,19,21}

Many studies on the preparation of CIS/ZnS core/shell NCs have indeed reported enhanced PL QYs (as high as \textasciitilde 80\%) and increased stability.\textsuperscript{11,15,16,25,35} However, spectral blue-shifts in both absorption and PL have been invariably observed after ZnS shelling of CIS NCs, in striking contrast with the small red-shift expected for type-I core/shell hetero-NCs due to leakage of the exciton wave function into the shell.\textsuperscript{29} This widening of the band gap of the NCs after ZnS shelling has, prior to onset of heteroepitaxial shell overgrowth,\textsuperscript{11,18,19} or increased stability.\textsuperscript{2,11} Leakage of the exciton wave function into the shell.\textsuperscript{29} This red-shift expected for type-I core/shell hetero-NCs due to exchange.\textsuperscript{37}

Figure 1. Although this is to a certain extent true for any semiconductor NC, the CIS shelling chemistry appears to be far more complex than that of other well-investigated systems, regardless of the precursors used, with the dominant process being determined by the nature of the precursors. Additive heteroepitaxial ZnS shell overgrowth becomes the dominant process only if reactive S- and Zn-precursors (S-ODE/OLAM and ZnI\textsubscript{2}) and high shelling reaction temperatures (210 °C) are used, although accompanied by a certain degree of hetero-interfacial alloying. Remarkably, we observe that residual acetate at the surface of the CIS seed NCs favors heteroepitaxial ZnS shell overgrowth, while depressing alloying, etching, and cation exchange to negligible levels. This yielded, for the first time, CIS/ZnS core/shell NCs displaying red-shifted absorption spectra. Finally, we discuss the mechanisms behind the chemical processes that take place during ZnS shelling reactions on CIS NCs, proposing a model that explains the impact of different physical-chemical variables on the balance between these processes, ultimately dictating the outcome of the shelling reaction. These insights can be used to design synthesis strategies to CIS/ZnS core/shell and alloy NCs with tailored elemental distribution profiles, allowing precise tuning of the optoelectronic properties of the resulting materials.

\section*{Experimental Section}

**Materials.** Copper(I) iodide (CuI, Sigma-Aldrich, 98\%), indium (III) acetate (In(Ac)\textsubscript{3}, Sigma-Aldrich, 99.99\%), 1-dodecanethiol (DDT, Sigma-Aldrich \textasciitilde 98\%), zinc stearate (ZnSt\textsubscript{2}, Sigma-Aldrich, 10–12\% Zn basis), zinc iodide (Sigma-Aldrich, 98+\%), 1-octadecene (ODE, Sigma-Aldrich, tech. 90\%), sulfur (Sigma-Aldrich 99.98\%), oleylamine (OLAM, Sigma-Aldrich, tech. 70\%), toluene (Sigma-Aldrich 99.8\%), methanol (Sigma-Aldrich 99.8\%), 1-butanol (Sigma-Aldrich 99.8\%), ethanol (Sigma-Aldrich \textasciitilde 98.8\%), acetone (VWR international, dried max 0.0075\% H\textsubscript{2}O), nitric acid (Sigma-Aldrich, 65\%). Prior to usage, ODE and OLAM were degassed for 2 h under vacuum at, respectively, 200 and 150 °C. All other chemicals were used as received.

**Chalcopyrite CuInS\textsubscript{2} NC Synthesis.** The CIS NCs were synthesized based on a protocol reported by De Trizio et al.\textsuperscript{14} Typically, CuI (0.8 mmol) and In(Ac)\textsubscript{3} (0.8 mmol) in 10 mL of DDT were mixed under inert atmosphere and degassed for 1 h at 80 °C. Under N\textsubscript{2} atmosphere, the temperature was raised to 210 °C, and after 40 min of reaction, the mixture was cooled down to room temperature. To investigate the effect of the washing protocol on the outcome of the shelling reactions, the NCs were washed by precipitation once with one of three different antisolvents: 1:1 methanol/butanol mixture, ethanol or acetone, the total volume of antisolvent being 10 mL. The precipitate was isolated by centrifugation and redispersed in 15 mL of toluene.

**ZnS Shelling Reactions.** ZnS shelling experiments were carried out by using a seeded injection technique in which a mixture of CIS NC seeds and a sulfur precursor was swiftly injected under stirring into a hot solution of the Zn-precursor in ODE at either 150 or 210 °C. After 2 h, the reaction was stopped by removing the heating source. The product NCs were washed by precipitation with isometric methanol/butanol, isolated by centrifugation, and redispersed in toluene. To study the influence of the shell precursor reactivity, a highly reactive and a less reactive precursor was used for both the cation and the anion precursors. For the Zn-precursor, ZnI\textsubscript{2}, and ZnI\textsubscript{3} (\Delta H = −208 kJ/mol)\textsuperscript{39} were chosen as, respectively, the least reactive and the most reactive precursor. For the shelling reaction, 0.32 mmol Zn-precursor (ZnSt\textsubscript{2}) or ZnI\textsubscript{2} was mixed with 4 mL of ODE and heated for 30 min at 100 °C before being further heated to the reaction temperature (150 or 210 °C). As highly reactive S-precursor, elemental sulfur dissolved in ODE (\Delta H = 0 kJ/mol)\textsuperscript{40} was used, while DDT (\Delta H = −328.1 kJ/mol)\textsuperscript{41} was chosen as the least-reactive precursor. Sulfur in ODE (0.1 M) was prepared by dissolving 160 mg of elemental sulfur in 50 mL of ODE. The mixture was heated to 200 °C and stirred until a clear solution was obtained. The CIS seed NCs...
(4 mL of a solution of purified NCs in toluene) were sedimented by centrifugation and, depending on the sulfur precursor, redispersed in either 2 mL of DDT and 2 mL of ODE or 3 mL of 0.1 M S-ODE and 1 mL of OLAM. OLAM was only added as ligand to the S-ODE solution since DDT could serve as both sulfur source and ligand. It has been shown that at high temperatures (<100 °C) elemental sulfur reacts with OLAM (and alkylamines in general) to form H₂S and a number of reactive sulfur-containing species (e.g., alklylthioanides, alkylthiylamines)⁴²,⁴³ which may be the actual reactive S-precursor in the reactions carried out in the present study.

**Optical Characterization.** Absorption spectra were recorded with a PerkinElmer lambda 950 UV–vis/NIR spectrophotometer. PL measurements were performed with an Edinburgh Instruments FLS920 spectrofluorimeter equipped with a 450 W Xe lamp, a double excitation monochromator, and emission monochromator. The signal was detected with a Hamamatsu R928 PMT detector or, when the emission was at 800 nm or longer, with an Acton research SpectraPro 300i CCD camera with optical fiber. The magnitude of the shift of the peak positions in the optical spectra was calculated by converting the spectra to eV scale, according to the method reported by Edger et al.⁴⁴ and comparing the peak positions of the product NCs to the peak position of the seed CIS NCs. The peak position of the absorption spectra was determined by taking the second derivative of the spectra.

**Electron Microscopy.** Transmission electron microscopy (TEM) images were recorded with a Tecnai microscope (Thermo Fisher Scientific, formerly FEI company), with an acceleration voltage of 100 or 200 kV. Samples for TEM imaging were prepared by drop-casting a toluene solution of NCs onto a carbon-coated copper TEM grid. Energy dispersive X-ray spectroscopy (EDS) measurements were performed with a PerkinElmer Optima 8300 instrument and the wavelengths used for calibration were as follows: λ₉₂₅ = 240.2 nm; λ₃₂₄ = 324.725 nm; λ₃₂₅ = 325.609 nm. For EDS measurements, a Gatan 694 CCD camera and an EDAX spectrometer, Energy dispersive X-ray spectroscopy (EDS) measurements were performed on a Tecnai microscope (Thermo Fisher Scientific, formerly FEI company), with an acceleration voltage of 100 or 200 kV. Samples for TEM imaging were prepared by drop-casting a toluene solution of NCs onto a carbon-coated copper TEM grid.

**RESULTS AND DISCUSSION.**

The shellion reaction procedure used in our study was the seeded injection method (see Experimental Section for details). This method was chosen to keep the CIS seed NCs as long as possible separated from heat and the precursors to minimize undesired reactions between the precursors and the CIS surface (see Figure 1). The use of trioctylphosphine (TOP) was also intentionally avoided, due to its strong binding affinity to Cu(I), which would enhance cation exchange processes by promoting the creation of Cu(I) vacancies.⁴⁵−⁴⁷ The chalcopyrite CIS NCs used as seeds have a trigonal pyramidal shape with an average height of ~2.5 nm (Figure 2A), and show a featureless absorption spectrum with the lowest energy absorption centered at 530 nm, and a broad PL peak with maximum at 676 nm (Figure 2B). These optical properties are in line with those typically reported for CIS NCs.¹−⁶

**Effect of Washing on Surface Chemistry.** Since most processes that can take place during ZnS shelling reactions involve the CIS NC surface (see Figure 1 above), residual precursors and ligands that are not completely removed by the washing procedure could have a large impact on the outcome of the shelling reaction. To investigate the impact of different washing procedures on the surface chemistry of CIS NCs, a batch of freshly synthesized CIS NCs was divided into four vials

![Figure 2.](image-url)
and each portion was washed differently. Either ethanol, methanol/butanol (1:1 volume mixture), or acetone was used as antisolvent, and one portion was not washed to serve as reference. The absorption and PL spectra of the different portions after one washing cycle are shown in Figure 3. The dashed vertical lines indicate the peak position of the absorption and PL spectra of the unwashed CIS NCs. The key spectral features, such as the peak positions in PL and absorption spectra and PL line width, are preserved, which indicate that the electronic transitions involved are not affected by the washing procedure. If the different washing treatments had led to significant changes in elemental composition, size, or shape, this would have affected the band gap of the NCs and the exact peak positions. The only observed change is that the PL QY of the NCs decreased after all three washing procedures (Figure 3). This is commonly observed for colloidal NCs and is induced by partial removal of capping ligands, thereby leaving unpassivated surface sites, which act as traps for the photogenerated carriers.29,48

Table 1. Cu/In Ratios of CIS NCs after Different Washing Procedures Determined with Three Different Techniques

<table>
<thead>
<tr>
<th>Method</th>
<th>EDS</th>
<th>ICP</th>
<th>XPS</th>
</tr>
</thead>
<tbody>
<tr>
<td>unwashed</td>
<td>0.45 ± 0.004</td>
<td>0.18 ± 0.01</td>
<td></td>
</tr>
<tr>
<td>ethanol</td>
<td>0.94 ± 0.006</td>
<td>0.63 ± 0.01</td>
<td>0.46 ± 0.03</td>
</tr>
<tr>
<td>methanol/butanol</td>
<td>0.59 ± 0.003</td>
<td>0.61 ± 0.01</td>
<td>0.49 ± 0.03</td>
</tr>
<tr>
<td>acetone</td>
<td>0.68 ± 0.002</td>
<td>0.59 ± 0.02</td>
<td>0.58 ± 0.04</td>
</tr>
</tbody>
</table>

It is important to note that although the standard deviations for the ratios obtained for each technique can be precisely determined (see Experimental Section), the uncertainties associated with the different techniques are hard to estimate, and stem from the inherent biases associated with each technique. EDS is an electron microscopy technique in which a region on the grid is selected where mostly NCs are present. The data obtained thus reflect the elemental composition of the NCs but have an inherent uncertainty since only a small fraction of the NC ensemble can be measured at any given time. To minimize this uncertainty, we measured relatively large areas and several different spots on the grid. XPS, on the other hand, probes a larger area of a drop-casted sample, with regions of different thickness and, in case of inefficient washing, likely also different concentrations of residual ligands and unreacted precursors. If one considers the fact that the escape depth of the photogenerated electrons is limited (see below), it becomes clear that XPS may be biased toward elements in the organic matrix surrounding the NCs, if this matrix is too thick. Finally, ICP will provide the bulk composition of the sample including ligands, NCs, unreacted precursors and residual complexes formed in situ. The limitations of the different techniques are particularly critical for the unwashed sample, which cannot be analyzed by ICP (see Experimental Section), and is particularly prone to a bias toward the unreacted precursors and ligands. For this reason, one cannot meaningfully compare the Cu/In ratios observed for the washed sample with those obtained for the washed samples, since they may be distorted by the presence of unreacted precursors. Nevertheless, small changes in the cation ratios as a result of the washing procedures cannot be completely excluded.

XPS is not only suitable to determine the relative concentrations of elements present in a sample, but also allows distinguishing between different chemical species of an element as the exact binding energy (BE) of electrons is determined by the oxidation state and chemical environment of the atom. For bulk materials, XPS is known as a surface sensitive technique due to the limited escape depth of the generated photoelectrons.49 However, the CIS NCs investigated here have sizes of ~2.5 nm, so the escape depth is sufficiently large to probe a substantial part of a NC, including the ligand layer. High-resolution XPS spectra of relevant elements are shown in Figure 4A–D (XPS survey spectra are shown in the Supporting Information, Figure S2, and high-resolution XPS spectra of iodide 3d are shown in Figure S3). The indium 3d5/2 peak is observed at a BE of 445.3 eV in the unwashed sample. In the washed samples, this peak shifts to lower energies (~444.9 eV, Figure 4A), which is in good agreement with BEs earlier in the washing procedure.

Figure 3. Absorption and PL spectra of unwashed CIS NCs (black) and CIS NCs that were washed with either ethanol (red), methanol/butanol 1:1 mixture (green), or acetone (blue). The dotted lines mark the absorption and PL maxima. The relative PL QYs are also given. The NCs are all from the same batch and differ only with respect to the washing procedure.
The higher BE measured in the unwashed sample may result from a contribution of unreacted indium acetate, resulting in a convolution of peaks (the In 3d5/2 peak of the similar complex indium acetylacetonate is found at a BE of 445.6 eV). However, the fwhm of the peaks (viz., 1.3 and 1.2 eV for the unwashed and washed samples, respectively) does not justify fitting more than one peak to account for different In-containing species.

The copper 2p3/2 peak lies around a BE of 932.5 eV and thus partially overlaps with the I 3p1/2 peak around 930.7 eV. The fitting of the Cu peaks was therefore corrected for the I-overlap (see Experimental Section for details). In the unwashed sample, the Cu 2p3/2 peak has a BE of 932.7 eV, whereas in the washed samples this peak shifts to lower energies (932.3–932.4 eV, Figure 4C). These BEs are again in good agreement with energies reported earlier for Cu(I) in CIS NCs. Because the resolution in the XPS spectra is 0.2 eV and the copper 2p3/2 peak in CuI is expected at a similar BE (932 eV), a convolution of signals cannot be excluded. The presence of Cu2+ in the samples is however excluded, as no satellite features around 942 eV associated with divalent copper have been observed (Figure 4C). It is interesting to note that the small shifts in BEs for Cu and In, between the unwashed and the washed samples, could also be ascribed to small compositional differences between the samples, as De Trizio et al. reported shifting of the Cu 2p3/2 peak from 932.3 to 932.6 eV and of the In 3d5/2 peak from 444.9 to 445.3 eV, upon decreasing the Cu/In ratio.14

The sulfur 2p3/2 peaks at BEs of 163.0 (163.2 for the unwashed sample) and 162.0 eV (Figure 4B) show that in all samples two sulfur species are present. To identify these sulfur species, we compare different literature studies on the interactions of alkanethiols with surfaces. First, thiols chemisorbed on a Cu(100) surface form a thiolate complex and have a sulfur 2p3/2 BE of 162 eV, while after thermal desorption of the alkyl chain the BE shifts to 161.3 eV for the remaining atomic sulfur. Second, on a InP(100) surface, the sulfur BE of alkanethiols is 162.4 eV, while for elemental sulfur at this same surface a BE of 161.6 eV has been reported. Third, crystal bound DDT molecules in Cu2S NCs give rise to S 2p3/2 signal at a BE of 161.8 eV (sulfur as part of the crystal lattice, i.e., three-fold coordinated to Cu(I) atoms), while in Cu2S NCs with surface bound DDT molecules peaks at both 161.8 and 162.2 eV are measured, which were ascribed to sulfur as part of the crystal structure and surface-bound thiolate, respectively. The trend in all these studies is clear: upon adsorption, chemisorption, and subsequent desorption of the alkyl chain, leaving sulfur behind on or in the crystal, the S 2p3/2 BE shifts to lower values. We therefore ascribe the higher BE of 163.0 eV in our study to thiolates at the surface of the NC. The slightly higher BE of 163.2 eV in the unwashed sample is ascribed to the presence of DDT molecules. The peak at lower BE (162.0 eV) is attributed to sulfur ions that are part of the CIS NC, in agreement with a recent study. As the fwhm of the sulfur related peaks is 1.2–1.4 eV, fitting additional peaks related to other sulfur-containing species, such as free DDT or didodecylsulfide molecules, is not justified, although a recent study by Reiss and co-workers reported the presence of these molecules.

Carbon species are detected in all samples and can only originate from thiolate molecules or residual acetate from the indium acetate precursor since the antisolvents used in the washing procedures have low boiling points and thus evaporate under the ultrahigh vacuum present in the XPS measurement chamber. In all samples, C 1s peaks are detected at BEs of 284.8, 286.8–287.1, and 285.8 eV, that we ascribe to the thiolate molecule. A fourth peak at a BE of 288.6 eV is clearly present in the ethanol washed sample (Figure 4D). In the other samples, this peak is quite weak and is fitted in a range of 289.1 to 288.1 eV (Figure 4D). C is BEs in this region are associated with esters or carboxylic groups and are in our case
likely that at least one localized carrier is involved in the as precursors (see Experimental Section for details). These washed with ethanol or methanol/butanol (i.e., relative thiolate content decreases most for the samples that are washed sample, and 9.5 for the ethanol washed sample). The four di temperatures (150 and 210°C) seed for shelling reactions at two different washing with ethanol leaves a signifi- coordinating molecules (both thiolates and acetate) and that methanol and butanol is the most e observations, we conclude that an isometric mixture of residual acetate in the CIS NC samples. The absorption spectra of the product NCs of all eight shelling reactions are shown in Figure 5. The spectra of the product NCs using the unwashed core NCs are clearly red-shifted with respect to the seed NCs, indicating a reduction of the quantum confinement. This observation can be explained by considering the excess of CIS precursors in the unwashed sample, which leads to further growth of the CIS NCs upon injection in the hot reaction mixture at both reaction temperatures (150 and 210 °C), employing ZnI2 and S-ODE/OLAM as precursors (see Experimental Section for details). These precursors were chosen because they are both very reactive, making it less likely that the monomer formation rates dominate the reaction kinetics, thereby allowing the impact of competing reactions involving the surface (i.e., heteroepitaxial growth, alloying, etching, cation exchange, see Figure 1 above) to be more evident. Optical spectroscopy was used as main technique to investigate the outcome of the seeded shelling reactions, as shifting of the absorption peaks gives a clear indication of changes in the effective core size resulting from alloying, etching, cation exchange (blue-shift), or heteroepitaxial shell overgrowth (red-shift). We chose to use absorption rather than PL for this analysis, as there is still a lively debate on the origin of the PL in CIS NCs.1,4,12,61,62 It is likely that at least one localized carrier is involved in the radiative decay pathway, but the nature of this localized state is as yet unclear.1,4,12,61,62 Therefore, one can only speculate on the impact of alloying, cation exchange or shell overgrowth on this localized state and the resulting PL spectral shifts. This difficulty is further aggravated by the fact that the PL QYs of bare CIS NCs are typically very low and drastically increase after shelling reactions, therefore making the comparison between the PL spectra prior to and after shelling reactions unreliable since the PL spectrum of the CIS seed NCs may not necessarily reflect the size, shape, and composition polydisper- sity of the ensemble. In contrast, the lowest energy absorption transition is only a function of the band edge positions, which are directly related to the degree of alloying, cation exchange, or shell overgrowth.

The absorption spectra of the product NCs of all eight shelling reactions are shown in Figure 5. The spectra of the four differently washed CIS NC samples were used as seeds for shelling reactions at two different reaction tempera- tures (150 and 210 °C), employing ZnI2 and S-ODE/OLAM as precursors (see Experimental Section for details). These precursors were chosen because they are both very reactive, making it less likely that the monomer formation rates dominate the reaction kinetics, thereby allowing the impact of competing reactions involving the surface (i.e., heteroepitaxial growth, alloying, etching, cation exchange, see Figure 1 above) to be more evident. Optical spectroscopy was used as main technique to investigate the outcome of the seeded shelling reactions, as shifting of the absorption peaks gives a clear indication of changes in the effective core size resulting from alloying, etching, cation exchange (blue-shift), or heteroepitaxial shell overgrowth (red-shift). We chose to use absorption rather than PL for this analysis, as there is still a lively debate on the origin of the PL in CIS NCs.1,4,12,61,62 It is likely that at least one localized carrier is involved in the
These product NCs also have a lower colloidal stability than those obtained from ethanol or methanol/butanol washed seeds, which resulted in the formation of aggregates that lead to a light scattering background in the absorption spectra of these samples (see 210 °C curve in Figure 5d). Interestingly, washing with methanol/butanol or acetone resulted in spectral blue-shifts at both 150 and 210 °C (78 and 43 meV, respectively, for methanol/butanol, and 43 and ∼50 meV, respectively, for acetone washed samples), while the use of ethanol washed CIS seed NCs leads to a blue-shift of 119 meV at 150 °C and a red-shift of 61 meV after reaction at 210 °C. The latter observation is particularly relevant since in this case the observed red-shift cannot be ascribed to growth of the seed CIS NCs, as CIS core-precursors were washed out. These results thus imply that residual acetate favors heteroepitaxial shell overgrowth at 210 °C, thereby leading to a spectral red-shift due to exciton leakage in the shell while promoting etching and/or alloying reactions at low reaction temperatures. Moreover, these observations highlight the crucial impact of the reaction temperature on the balance between the different chemical processes taking place during the ZnS shelling reaction, thereby influencing not only the extent of the blue-shift induced by alloying, cation exchange, or etching reactions, but also the heteroepitaxial shell overgrowth. These inferences will be confirmed below by structural investigations. Clearly, both the reaction temperature and the surface chemistry play a crucial role in determining the balance between the different chemical processes that can take place during the shelling reaction. The interplay between these variables will be discussed in more detail in the mechanism section below after we address the impact of the precursor reactivity and investigate the atomic structure of representative CIS/ZnS product NCs.

It should be noted that the impact of the ZnS shelling on the PL QYs of the CIS NCs is strongly dependent on the washing procedure, the largest enhancement being observed for unwashed seed NCs (factor 2.7 at 150 °C and 3.5 at 210 °C), followed by ethanol washed and methanol/butanol washed seed NCs (factor 1 to 1.6 at 150 °C and 1.2 to 2.7 at 210 °C, depending on the batch of seeds). The acetone washed seed NCs actually result in product NCs with lower PLQYs than the initial ones (reduction by a factor 3). These differences are intriguing and suggest that the state of the surface of the seed NCs is largely responsible for the quality of the CIS/ZnS heterointerface formed during the ZnS shelling reaction. It is also remarkable that the PL QY enhancements observed in the present work after shelling are more modest than those previously reported in the literature using similarly sized chalcopyrite CIS NCs as cores (factor 10−15).11,12 It is as yet unclear whether these differences are due to the different growth methods used (viz, seeded injection in the present work and one-pot slow addition in refs 11 and 12) or to other variables such as the nature of the shell precursors and ligands present in the reactions. Sheding light on these questions, however, is beyond the scope of this work and will be the subject of future follow-up work.

Precursor Reactivity. The results discussed above show that both the reaction temperature and the presence of residual acetate and thioclate in the CIS seed NCs have a dramatic impact on the outcome of the shelling reaction. Another crucial variable is the precursor reactivity, which determines the monomer formation rate and therefore directly impacts on the kinetics of both the heteroepitaxial growth and the homogeneous secondary nucleation (Figure 1).29,36,63−67 Further, the precursor reactivity also affects the rates of undesired parallel processes such as cation exchange, etching, and alloying (Figure 1). Methanol/butanol washed CIS NCs were used as seeds to investigate the influence of the precursor reactivity since this washing was the most effective in removing coordinating molecules (both acetate and thioclate, see above), therefore minimizing possible synergistic or antagonist interactions with the precursors, allowing the isolated effect of the precursor reactivity to be more clearly observed. It should be noted that the precursor reactivity is determined by both its intrinsic stability and the reaction temperature. Therefore, also in this case two different shelling reaction temperatures were used. The outcome of the shelling reactions was followed not only by the extent of the spectral shifts observed for the product NCs, but also by the change in their elemental composition with respect to the CIS seed NCs (Figure 6 and Supporting Information Figure S5). Taken together, these variables allow the dominant chemical processes acting on the CIS seed NCs during the ZnS shelling reactions to be identified.

The data corresponding to the product NCs obtained from the reaction using ZnSt2 and S-ODE as precursors is not included in Figure 6, since in this case large 2D nanoplatelets and nanosheets were obtained at both reaction temperatures (Figure S6). This intriguing outcome has been studied in detail in previous work by our group10 and shown to be induced by
A precursor: an unreactive Zn-precursor, such as ZnSt$_2$, is unable of the etching is directly related to the reactivity of the Zn-precursor, capable of quickly extracting In$^{3+}$ (and to a lesser extent Cu$^+$) from CIS NCs, thereby leading to pronounced etching, which is characterized by large blue-shifts accompanied by large content, accompanied by unchanged (or slightly increased) In/Cu ratios (Figure 6). Interestingly, decreasing the reaction temperature to 150 $^\circ$C has a higher impact on the reactivity of ZnI$_2$ (and thereby on the $[\text{ZnS}]$ formation rate) than on the reactivity of S-ODE/OLAM, thereby shifting the balance toward the NC etching reaction, resulting in larger blue-shifts, lower Zn-content, and larger decrease in the In/Cu ratio, with respect to the reaction carried out at 210 $^\circ$C.

The data in Figure 6 show that shelling reactions at 150 $^\circ$C (blue data points) invariably lead to larger blue-shifts, lower Zn/Cu ratios, and larger reductions in In/Cu with respect to the same reaction carried out at 210 $^\circ$C. This shows that lower reaction temperatures favor etching, alloying, and cation exchange. Zn$^{2+}$ for Cu$^+$ or In$^{3+}$ CE is favored by using more stable precursors, such as Zn(St)$_2$ and DDT, since in this case the precursor to monomer conversion rate will be slow, leaving the Zn-precursor available to bind to the NC surface and undergo cation exchange. This process is further favored in the presence of DDT because this molecule is unable to induce extensive etching, being already the capping ligand shell on CIS NCs synthesized with DDT as S-source, thereby leaving a relatively stable surface for adsorption of the Zn-precursor. The prevalence of CE under these conditions is clearly evidenced by the observation that the shelling reaction using Zn(St)$_2$ and DDT at 210 $^\circ$C leads to the largest blue-shift and the highest Zn-content. Moreover, replacing Zn(St)$_2$ by the more reactive ZnI$_2$ leads to smaller blue-shifts and lower Zn-contents. This is in line with the expected trend in the CE efficiencies since the driving force for the CE reaction is the energy gain upon replacing the native cation by the guest cation. Considering that In$^{3+}$ is a hard Lewis acid, the formation of In(III) stearate is more favorable than that of In(III) iodide, as stearate is a harder Lewis base and stronger ligand than iodide. This also explains why Zn$^{2+}$ for In$^{3+}$ CE is more efficient than that of Zn$^{2+}$ for Cu$^+$. It is thus clear that the outcome of a ZnS shelling reaction on CIS NCs is determined by a delicate balance between the rates of all the chemical processes depicted in Figure 1 and that this balance is highly sensitive not only to the reaction temperature and the nature of the Zn- and S-precursors, but also to the presence of residual coordinating molecules at the surface of the NCs. The mechanism behind this complex interplay of physical-chemical processes will be discussed in more detail below after the atomistic structure characterization of two representative examples of product NCs.

**Structural Characterization.** The observations described above suggest that the use of reactive S- and Zn-precursors (ZnI$_2$ and S-ODE/OLAM) at high temperatures (210 $^\circ$C) in
the presence of acetate leads to heteroepitaxial ZnS shell overgrowth on CIS NCs, with negligible contribution of undesired competing processes such as CE, alloying, and etching, while low reaction temperature (150 °C) or unreactive precursors shift the balance toward the competing processes. To verify whether this is indeed the case, the structure and composition of the seed and product NCs were investigated with transmission electron microscopy. Figure 7A and C show high-resolution (HR) HAADF-STEM images of CIS NCs washed with ethanol and used as seeds for a shelling reaction with S-ODE/OLAM and ZnI₂ at 150 or 210 °C, respectively. In Figure 7B and D, HR-HAADF-STEM images of product NCs obtained from these reactions are shown. The absorption spectrum of the product NCs obtained at 150 °C is blue-shifted by 40 meV ([In/Cu]_{product/seed} = 0.5), while that of the product NCs obtained at 210 °C is red-shifted by 200 meV ([In/Cu]_{product/seed} = 1.0). It is immediately clear that the shape of the seed NCs (trigonal pyramid) is essentially preserved in the product NCs obtained at 150 °C, despite a small growth (base increases from 2.7 ± 0.3 nm to 3.2 ± 0.7 nm and height increases from 2.5 ± 0.3 nm to 3.0 ± 0.4 nm). In contrast, the product NCs obtained at 210 °C have a different shape and a significantly larger size (average diameter of 4.7 ± 0.8 nm) with respect to the seed NCs (trigonal pyramids with an average base of 2.5 ± 0.3 nm and an average height of 2.3 ± 0.3 nm). Moreover, the NCs obtained at 150 °C (Figure 7B) show a uniform intensity over the whole NC, indicating that they consist of (Cu,In,Zn)S₂ alloys. This is consistent with the observation of blue-shifted spectra after the reaction, as discussed above. In contrast, the NCs obtained at 210 °C (Figure 7D) show increased intensity in the center of the particle, suggesting that either the heavier elements are mostly located in the center or an increased thickness is present at the center of the NCs. Since the intensity in HAADF-STEM images scales with both the projected thickness and the atomic number of the present elements, such 2D images cannot unambiguously confirm core–shell architectures. Therefore, electron tomography was applied to study the internal structure of the CIS/ZnS product NCs. Figure 7E shows the 3D visualization of a CIS/ZnS core/shell NC investigated during the electron tomography experiment. In Figure 7F and G, 2D slices through the 3D reconstruction are shown, with indication of the position of the slices in Figure 7E. In Figure 7G, a more intense region (red color) is detected, compared to the orthoslice in Figure 7F. Since indium has a much larger atomic mass (Z_{In} = 49) than the other elements present (Z_{Cu} = 29 and Z_{Zn} = 30), its presence in the core of the NC is confirmed, while the surface is richer in the lighter elements copper and/or zinc. From these tomography results, we conclude that the ZnS shell grows anisotropically. This is similar to previous observations on wurtzite CdSe/CdS dot core/rad shell nanorods and 36,46,70–72 and nanobullets 36 and wurtzite graded AgInSe₂/ZnSe core/shell NCs, 73 suggesting that the facets of the trigonal pyramidal shaped CIS seed NCs are not equivalent. However, the CIS seed NCs and the product NCs (both the alloy NCs obtained at 150 °C and the CIS/ZnS core/shell NCs obtained at 210 °C) have the chalcopryite structure (Supporting Information, Figure S7), which, in contrast with the hexagonal wurtzite structure, does not possess polar facets. Therefore, the anisotropy observed in the heteroepitaxial shell overgrowth at 210 °C may be attributed to a different termination of the pyramid base with respect to the sides, either due to a different facet composition or due to selective adhesion of acetate, or a combination of both.

**Mechanism.** The results discussed above provide valuable insights on the chemical processes that take place during ZnS shelling reactions on CIS NCs (Figure 1), allowing us to understand the mechanisms behind them and the factors that determine the balance between these processes, ultimately dictating the outcome of the shelling reaction and the optoelectronic properties of the product NCs. These processes can be roughly divided into three categories: (i) reactions taking place in solution, such as the precursor to [ZnS] monomer conversion, homogeneous nucleation of ZnS NCs, and reactions involving the added precursors or ligands forming new species in situ (e.g., reaction between elemental sulfur and alkylamines forming H₂S, alkylthioamides, dialkylamidines); (ii) reactions with the surface (i.e., reactions that alter the chemical composition of the NC facets), such as etching, cation exchange, and alloying; and (iii) reactions on the surface, such as additive heteroepitaxial shell overgrowth and ligand adsorption or desorption. These reactions may occur in parallel, competing with each other for the limited supply of precursors, monomers, or available surface sites (e.g., etching, cation exchange, and alloying compete with heteropitaxial shell growth), or sequentially, in a concerted manner (e.g., monomer formation precedes both heteroepitaxial shell growth and homogeneous nucleation). Therefore, the outcome of the reaction depends on a complex interplay between a number of inherently linked elementary kinetic steps.

The precursor to monomer conversion has been shown to be the rate-limiting step in the formation of a variety of binary metal chalcogenides (viz., CdX, PbX, Cu₂S, X = S, Se), 29,36,63 and is likely also rate-limiting in heteroepitaxial shell overgrowth reactions, 29,36,67 since recent work has suggested that heteroepitaxial shells grow by incorporation of monomer units, rather than by layer-by-layer adsorption of atomic species, 29,67 even when the SILAR strategy is employed. 38 The adsorption of atomic species or precursors is in fact more likely to lead to etching, since the adsorbed species may leave again after reacting with the surface site, thereby carrying a metal or chalcogen atom with itself, before more precursors can join in to form a sufficiently stable adlayer. Etching can also be promoted by ligands or reactive species formed in situ. Alternatively, the adsorbed metal precursor may lead to cation exchange (CE). Nanoscale CE has been extensively investigated in recent years as a postsynthetic strategy to control the composition of NCs, and is thus well-understood. 14 The CE process itself is essentially a surface reaction, which would be self-limited in the absence of cation diffusion in the NC, stopping as soon as all surface native cations had been exchanged by guest cations. 38 Therefore, solid state diffusion fluxes must be set in motion to allow the reaction to proceed. Depending on the solid state diffusion rates of the guest and host cations, the CE process can lead to either superseded shell ingrowth (e.g., PbSe/CdSe, ZnSe/CdSe, Zn₂Se/CdSe, 37–39,68 or alloy NCs (graded or homogeneous, e.g., (Zn,Cd)Se, 36 CuInS₂ from Cu₂–S, 36 (Cu,In,Zn)S₂, 36,14 CsPbBr₃:M with M = Sn²⁺, Cd²⁺, Zn²⁺. 74). Diffusion may lead to alloying, even in the absence of cation exchange since Zn²⁺ cations from [ZnS] monomers deposited at the surface may diffuse in the CIS NCs if a sufficiently stable ZnS heteroepitaxial monolayer is not formed fast enough. This is expected because the binding strength of a Zn²⁺ in an isolated [ZnS] adsorbate is much weaker than that in
a fully grown ZnS shell due to its lower coordination number, resulting in much lower activation energies for diffusion.

Each of the processes described above has its own activation energy, and therefore depends differently on temperature. Consequently, the reaction temperature has a dramatic impact on the balance between the rates of the different processes, and can thus be used as a sensitive parameter to tailor the product of the shell growth. The observations discussed above show that low temperatures (150 °C) favor etching, alloying, and cation exchange, leading to pronounced spectral blue-shifts. This can be rationalized by considering that at this temperature the [ZnS] monomer formation rates are too low, and therefore, heteroepitaxial overgrowth by monomer addition is unable to outcompete the other processes. The unreacted shell precursors are then available to undergo direct reactions with accessible sites at the surface of the CIS NCs, reducing the monomer formation rates even further. The dominant process under these conditions will then depend on the nature of the precursors. Reactive S-precursors (S-ODE/OLAM) combined with unreactive Zn-precursors (ZnSt₂) lead primarily to etching, resulting in product NCs with large spectral blue-shifts (>150 meV) and low Zn-content, while the use of reactive S- and Zn-precursors (S-ODE/OLAM and ZnI₂) results in a combination of etching and [ZnS] deposition followed by interdiffusion, yielding product (CIS/ZnS) alloy NCs with intermediate blue-shifts (50–150 meV), high Zn content, and decreased In/Cu ratio with respect to the seed NCs.

Conversely, unreactive S-precursors (DDT) favor cation exchange involving the Zn-precursors, which is most efficient when the Zn-precursor is also less reactive (ZnSt₂). As discussed above, this is due to the fact that the driving force for CE reactions is the energy gain upon replacement of the native cation by the guest cation, which is larger if the Zn²⁺ precipitation for CE reactions is the energy gain upon replacement of the Zn-precursor by the guest S-precursor. 

Considering that the magnitude of the spectral blue-shift would be accompanied by incorporation of both Cu and In. Considering that the magnitude of the spectral blue-shift depends on the nature of the Zn interdiffusion and alloying (and on the extent of prior etching of the NC seeds), one can easily see that differences in the reaction temperatures and reaction times used in different studies will also add to the spread in the observed blue-shifts.

It should be noted that homogeneous nucleation of ZnS NCs was not observed in any of the experiments discussed above. This can be attributed to the fact that the activation energies for homogeneous nucleation are even higher than those for heterogeneous nucleation and heteroepitaxial overgrowth. Therefore, homogeneous nucleation would only be significant if the monomer formation rates would exceed the heteroepitaxial growth rates or if the NC surface would not be accessible to the [ZnS] monomers due to a too dense and strongly bound ligand layer, which was clearly not the case under the conditions prevalent in our study.

The insights obtained in our work can be used to design synthesis strategies to CIS/ZnS core/shell and alloy NCs and are also useful to rationalize the seemingly disparate results presented in the literature, with reported spectral blue-shifts ranging from 60 to 340 meV. The diversity of the trends reported in literature for ZnS shell growth reactions on CIS NCs can be understood in light of the mechanisms proposed in our study and can be ascribed to the wide variety of reaction conditions used. Most studies used the dropwise addition of relatively unreactive shell precursors (typically ZnSt₂ and TOP-S) to a hot (T ≥ 200 °C) solution containing CIS NCs. Such conditions would indisputably favor alloying since the [ZnS] monomer formation rates will be slow and the surface mobility will be high. The presence of TOP in many studies will also favor cation exchange and etching because TOP (a soft Lewis base with absolute hardness η ≈ 6 eV), has a strong affinity for Cu⁺ (a soft Lewis acid with η = 6.28 eV) and also for sulfur, thereby being able to extract both from the CIS NCs. Several studies have also used the crude reaction mixture (so unwashed CIS seed NCs), thereby adding residual Cu- and In-precursors to the shell growth reaction mixture. This should facilitate the formation of (gradient) alloy NCs even further, since the deposition and interdiffusion of [ZnS] monomers would be accompanied by incorporation of both Cu and In. Considering that the magnitude of the spectral blue-shift depends on the extent of the Zn interdiffusion and alloying (and on the extent of prior etching of the NC seeds), one can easily see that differences in the reaction temperatures and reaction times used in different studies will also add to the spread in the observed blue-shifts.
Interestingly, the crystal structure of the CIS seed NCs is not a relevant parameter from this viewpoint since a recent study by our group has shown that the spectral blue-shifts observed after ZnS shelling of wurtzite CIS NCs by dropwise addition of Zn- and S-precursors (viz., 50–150 meV, with unchanged In/Cu ratio) are analogous to those observed for chalcopyrite CIS NCs under similar conditions. However, under seeded growth conditions, the crystal structure of the CIS seed NCs has a dramatic impact on the morphology of the product CIS/ZnS hetero-NCs, leading to CIS/ZnS dot core/rod shell hetero-nanorods, if wurtzite CIS NCs are used as seeds and the [ZnS] monomer concentration is allowed to build-up by delaying the injection of the seeds with respect to the injection of the Zn- and S-precursors. This recent study clearly demonstrates the crucial role of the monomer formation rates in determining the outcome of shelling reactions, in agreement with the mechanism proposed above. There are also studies that used conditions that favor cation exchange (e.g., addition of only Zn-precursors), resulting in (CIS, ZnS) alloy NCs with high Zn-content and consequently very large blue-shifts (340 meV). In this context, the recent study by Woods and co-workers on the preparation of CIS/ZnSe alloyed core/shell NCs by ZnI2 for In3+ and Cu+ cation exchange is particularly interesting. The authors demonstrated that the thickness of the ingrown superseded shell was determined by the reaction temperature (lower temperatures yielding thinner shells), concluding that the diffusion of Zn2+ cations into successive atomic monolayers of the seed NCs is a thermally activated process in which the activation energy increases in a depth-dependent fashion. Similar observations have been reported for ZnSe/CdSe core/shell and graded alloy NCs demonstrating that the reaction temperature is a sensitive parameter to tailor the composition and the elemental distribution profile of semiconductor hetero-NCs from core/shell to homogeneous alloy NCs through graded alloys, thereby tuning the carrier localization regime and the optoelectronic properties of the resulting materials.

CONCLUSIONS
The outcome of ZnS shelling reactions on CIS NCs is determined by a delicate balance between several chemical processes that take place both sequentially and in parallel, competing with each other for the limited supply of shell precursors and available surface sites. In this study, we investigated the impact of the precursor reactivity, reaction temperature, and surface chemistry on the outcome of the shelling reaction using a seeded growth strategy. We demonstrate that low reaction temperatures (150 °C) favor etching, cation exchange, and alloying, regardless of the precursors used. Under these conditions, the dominant process is determined by the nature of the precursors used. Reactive S-precursors (S-ODE/OLAM) result primarily in etching, while the use of reactive S- and Zn-precursors (S-ODE/OLAM and ZnI2) results in a combination of etching and [ZnS] deposition followed by alloying. Unreactive S-precursors (DDT) favor cation exchange involving the Zn-precursors, which is most efficient with more stable Zn-precursors.

High reaction temperatures (210 °C) and less reactive precursors also favor cation exchange followed by alloying. Heteroepitaxial ZnS shell overgrowth only becomes dominant if reactive S- and Zn-precursors and high reaction temperatures are used because these conditions shift the balance toward [ZnS] monomer formation. Nevertheless, a certain degree of [ZnS] interdiffusion and heterointerfacial alloying still occurs, yielding product NCs, which still show spectral blue-shifts with respect to the seed NCs.

Remarkably, the presence of residual acetate at the surface of the CIS seed NCs is shown to depress etching, cation exchange, and alloying to negligible levels, while facilitating heteroepitaxial shell overgrowth, yielding for the first time CIS/ZnS core/shell NCs displaying red-shifted absorption spectra, in agreement with the spectral shifts expected for a type-I band alignment. This finding highlights the crucial importance of the surface chemistry of the CIS NCs and of the washing procedures used to purify the seed NCs prior to the shelling reaction since residual acetate is only present in samples washed with ethanol. The insights provided by this work pave the way toward the design of improved synthesis strategies to CIS/ZnS core/shell and alloy NCs with tailored elemental distribution profiles, allowing precise tuning of the carrier localization regime and the optoelectronic properties of the resulting materials.

ASSOCIATED CONTENT
Supporting Information
The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.chemmate.8b00477.

Representative EDS spectrum of CIS core NCs, XPS spectra, additional HR-HAADF-STEM images of CIS core NCs, alloy and core/shell product NCs, optical spectra of CIS core NCs and product NCs after reaction at different temperatures with different precursors, TEM images of product NCs after reaction with zinc stearate, sulfur, and oleylamine (PDF) 46, 50, 52, 53, 57-59

AUTHOR INFORMATION
Corresponding Author
*E-mail: c.demello-donega@uu.nl.

ORCID
Ward van der Stam: 0000-0001-8155-5400
Jan P. Hofmann: 0000-0002-5765-1096
Sara Bals: 0000-0002-4249-8017
Celso de Mello Donega: 0000-0002-4403-3627

Present Address
1W.v.d.S: Optoelectronic Materials Section, Faculty of Applied Sciences, Delft University of Technology, van der Maasweg 9, 2629 HZ Delft, The Netherlands.

Notes
The authors declare no competing financial interest.

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